

Appendix A

List of Symbols

For convenience, many symbols frequently used in the text are listed below alphabetically. Some of the symbols may play a dual role, in which case the meaning should be obvious from the context. A number in parentheses refers to an Eq. number (or the text or footnote near that Eq. number) in which the symbol is defined or appears for the first time; in a few cases a page number, Table number, Sect. number, footnote (fn.) number, or Fig. number is given instead. Quantum number is abbreviated as q. n.

Quantum numbers (atomic):

Symbol	Description	Eq. No.
$j = l \pm s$	Inner q. n. of an electron	(3.13)
$J = L + S$	Inner (total angular momentum) q. n. of a state	(5.51b)
l	Azimuthal (orbital) q. n. of an electron	(3.1) and (3.2)
$L = \sum l$	Azimuthal (orbital) q. n. of a state	(5.51c)
m_l	Magnetic orbital q. n. of an electron	(3.1)
m_s	Magnetic spin q. n. of an electron	(3.1)
M	Total magnetic q. n.	(5.51d)
n	Principal q. n. of an electron	(3.1)
s	Spin q. n. of an electron	(3.13)
$S = \sum s$	Spin q. n. of a state	(5.51e)
κ	Relativistic azimuthal q. n. of an electron	(3.15a) and (3.15b)
μ	Relativistic magnetic q. n. of an electron	(3.18a) and (3.18b)

Quantum numbers (molecular)¹:

Symbol	Description	Eq. No.
I	Nuclear spin q. n.	p. 62
J	Rotational q. n.	(3.63)
K	Total angular momentum q. n. apart from spin	(3.80)
L	Electronic orbital angular momentum q. n.	p. 61
S	Electron spin angular momentum q. n.	pp. 59–62
ν	Vibrational q. n.	(3.62)
Λ	q. n. projection of L on internuclear axis	pp. 59–62
Σ	q. n. for projection of S on internuclear axis	pp. 59–62
$\Omega = \Lambda + \Sigma$	Total angular momentum q. n. about internuclear axis	p. 59

General list of symbols:

Symbol	Description	Eq. No.
A	Atomic mass	
A	Conversion coefficient for extinction, absorption, and scattering	(2.56)
A_i	IR molar absorption lengths	(5.116)
$A_i(r)$	Exchange Coulomb interaction	(3.3) and (3.4a)
A_ν	Rotational constant for vibrational level (for diatomics the subscript ν can be deleted)	p. 61 and (3.94)
A_{ul}	Radiation Einstein coefficient	(5.56a) and (5.57a)
\mathbf{A}	Vector potential for electromagnetic field	(5.48)
\mathcal{A}	Spin coupling constant for a multiplet	(3.78)
\mathcal{A}_ν	Albedo at frequency ν	(2.8)
a_o	Radius of first Bohr orbit of hydrogen	p. 30
a_i^s	Scattering coefficient	(5.267a) and (5.269b)
\hat{a}_i	Activity of chemical species i	(4.74) and (4.76)
$a_S = 4\sigma/c$	Stefan's radiation energy density constant	(2.48)
$B(T)$	Planck radiance at temperature T	(2.36)
B_e	Rotational constant for equilibrium position	(3.64b)
$B_{ij}(r)$	Exchange Coulomb interaction	(3.3) and (3.4b)
B_{ul}	Radiation Einstein coefficient for radiation energy density	(5.56a) and (5.57b)
B_ν	Rotational constant for vibrational level	(3.64a) and (3.94)
$B_\nu(T)$	Planck photon frequency distribution function at frequency ν and temperature T	(2.4)

¹Herzberg (1950, Chap. V) includes helpful vector diagrams illustrating various cases of coupling of electronic and rotational motions.

B	Magnetic induction	(5.2)
b_i^s	Scattering coefficient	(5.267b) and (5.269b)
b	Ray's asymmetry parameter	(3.94)
C_{ij}	Collisional rate between states i and j	(2.1)
C_v	Rotational constant for vibrational level	(3.94)
c	Speed of light	
$c^{(k)}$	Gaunt angular coefficient	(3.8) and (3.9)
c_{ik}	Radiation scattering tensor	(5.210)
$D(\rho, T)$	Photon diffusion coefficient at density ρ and temperature T	(2.49)
$D^a(u)$	Reduced absorption coefficient without stimulated emission	(2.55b)
$D^{a'}(u)$	Reduced absorption coefficient with stimulated emission	(2.55a)
D_e	Width of Gauss distribution at e^{-1} of maximum	(7.15a)
$D^s(u)$	Reduced scattering coefficient	(6.60)
$D^{\mathcal{T}}(u)$	Reduced transport extinction coefficient	(2.55a) and (2.55b)
$D^{s\mathcal{T}}(u)$	Reduced transport scattering coefficient	(2.55a)
D_n	Binding energy of molecule	(3.65)
D	Electric displacement	(5.1)
$D(\theta, \phi, \chi)$	Dyad relating space-fixed axes to molecule-fixed axes in terms of Eulerian angles	(5.99)
d_i	Spin-orbit angular interaction of an electron	(3.12)
E	Total energy of molecule, atom, or ion	(3.59) and (4.1)
E_o	Continuum lowering energy	(4.2) and (4.6)
E_f	Interaction energy with free electrons	(4.1)
E_i	Binding energy of an electron in level i	(3.3), (3.15a), and (3.15b)
\overline{E}_i	Mean binding energy of electron in level i	(3.47)
$E_i^{(n)}$	Binding energy of level i in neutral atom	(3.47)
E_{ij}	Electron-electron interaction energy	(3.46)
$E_{ij}^{(i)}$	Electron-electron interaction energy in an ion	(3.48)
$E_{ij}^{(n)}$	Electron-electron interaction energy in a neutral atom	(3.50)
E_J	Rotational energy in a molecule	(3.61a)
E_{kin}	Mean kinetic energy of ion and electrons	(4.34)
E_n	Electronic energy in a molecule	(3.61a)
E_{pot}	Mean potential energy	(4.33)
E_v	Vibrational energy in molecule	(3.61a)
E_{vJ}	Sum of vibrational and rotational energy in molecule	(3.67)
E	Electric field	(5.2)
$ E_o $	Magnitude of electric field vector E_o (in contrast to continuum lowering energy; see E_o)	(5.260a)
\mathcal{E}_i	Energy of excitation above ground state of ionization stage I	(4.59) and (4.60)
\mathcal{E}_v	Vibrational energy eigenvalue	p. 161 fn. 21
e	Charge of electron	
$\text{erf}(x)$	Error function	(6.4)

(continued)

(continued)

Symbol	Description	Eq. No.
e_v	Polarization vector	(5.55)
F	Helmholtz free energy	(4.21)
$F^+(\rho, T)$	Total outflowing flux density (exitance) at density ρ and temperature T	(2.34)
$F^-(\rho, T)$	Total inflowing flux density (irradiance) at density ρ and temperature T [in contrast see $F^+(\rho, T)$]	
$F^{(k)}$	Slater integral	(3.8b)
$F_i(r)$	r times the small Dirac wave function	(3.15a)
F_J	Rotational term value in molecule	(3.61b)
F_v^+	Outflowing component of flux (exitance) at frequency ν	(2.14)
F_v^-	Inflowing component of flux (irradiance) at frequency ν	(2.15)
$F(\mathbf{k})$	Form factor	(5.204)
$F_{\text{coh}}(\nu)$	Coherent form factor	(5.214)
$F_{\text{inc}}(\nu)$	Incoherent form factor	(5.216)
$\mathbf{F}(\mathbf{r})$	Conductive heat flow	(9.1a)
\mathbf{F}	Total flux	(2.46)
\mathbf{F}_ν	Net flux at frequency ν	(2.13)
$f(\eta, \delta)$	Correction factor for electron correlation	(6.51)
$f(v)dv$	Maxwell velocity distribution function	(6.14)
$f_e(\mathbf{x}, \mathbf{p})$	Electron distribution function	(9.17)
$f_i(\mathbf{x}, \mathbf{p})$	Ion distribution function	(9.22)
f_{lu}	Oscillator strength	(5.59)
f_S	Surface geometry coefficient	(2.30)
G	Gibbs free energy	(4.67)
$G_i(r)$	r times the large Dirac wave function	(3.15b)
$G^{(k)}$	Slater integral	(3.8)
G_ν	Vibrational term value in molecule	(3.61b)
$G_{L_i S_i}^{LS}$	Racah fractional parentage coefficient	(5.85b)
$\mathcal{G}(v)$	Spectral function of a binary complex of a supermolecular transition in a rarefied gas, such as $\nu_1 j_1 \nu_2 j_2 \rightarrow \nu'_1 j'_1 \nu'_2 j'_2$, thermodynamically averaged at a given temperature [identical to $\check{g}(v; T)$ for atom-atom pairs].	(8.32) and (8.33)
g	Bispinor	(3.19)
g_i	Statistical weight of state or level i	(2.2) and (5.61)
$g_{\text{bf},i}$	Bound-free Gaunt factor for an electron in level i	(5.86)
g_{ff}	Free-free Gaunt factor	(5.95) and (5.97)
$\check{g}(v; T)$	Spectral function [weighted sum of $\mathcal{G}(v)$]	(8.32)
$\check{g}(v; \nu\nu', T)$	Spectral function [see also $\phi(v)$]	(8.39)
H	Enthalpy	(4.67)
H	Total Hamiltonian ($H = H_0 + H'$)	(5.163) and (5.179)
H_0	Hamiltonian without radiation interaction	(5.150)
H'	Hamiltonian for radiation interaction	(5.46) and (5.48)

H_J	Hamiltonian for rotation	(3.77)
H_{nv}	Hamiltonian for electronic vibration	(3.77)
$H(\beta)$	Holtmark distribution	(7.72)
$H_n(x)$	Hermite polynomial	(5.131), Appendix C
H	Magnetic field	(5.4)
\mathcal{H}	Voigt function	(7.18)
\mathcal{H}'	Total Hamiltonian (including higher order terms)	(5.45)
h	Planck constant	
$h_l^{(1)}(x)$	Spherical Hankel functions of the first kind	(5.266), Appendix C p. 250 fn. 73
$h_l^{(2)}(x)$	Spherical Hankel functions of the second kind	(5.266), Appendix C p. 250 fn. 73
I	Irradiance (radiative flux), magnitude of Poynting vector	(5.35a)
I_A	Moment of inertia about the A axis (for diatomics it is the moment of inertia of the electrons about the internuclear axis)	p. 61
$I_k(\eta)$	Fermi integral of order k	(4.13)
I_ν	Spectral radiance at frequency ν	(2.3)
\mathcal{I}	Action	(3.67)
J	Total zeroth angular moment of radiance	(2.52)
J_ν	Zeroth angular moment of spectral radiance at frequency ν	(2.18)
\mathbf{J}_e	Current density of free electrons	(5.4)
$j_n(x)$	Spherical Bessel function of the first kind	(5.242), Appendix C
j_ν	Emission coefficient per unit mass at frequency ν	(2.11)
$K_{\hat{a}}$	Equilibrium constant for activity	(4.78)
K_c	Equilibrium constant for concentration	(4.80)
K_p	Equilibrium constant for pressure p	(4.79)
$K_n(x)$	Modified Bessel function of the second kind and order n	(6.30) and (7.31), Appendix C
K_ν	Second angular moment of spectral radiance at frequency ν	(2.23)
\mathbf{K}	Angular momentum operator	(3.80) and (3.93)
k	Boltzmann constant	
\mathbf{k}	Complex wave vector	(5.15) and (5.17)
\mathbf{k}'	Real part of wave vector	(5.14)
\mathbf{k}''	Imaginary part of wave vector	(5.14)
L	Total angular momentum	p. 33
L	Lorentz factor	(9.38)
L_{vW}	Vleck–Weisskopf impact line profile	(7.97)
L_{ZN}	Zhevakin–Naumov line profile	(7.98)
L_ν	Line shape function	(5.60a)
\mathcal{L}_ν	Line shape function (Lorentz and resonance)	(5.63) and (5.208)
l_p	Length of radiation path for a given optical depth at a given pressure	(11.20)
M	Molecular (also atomic) mass	(2.56)

(continued)

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Symbol	Description	Eq. No.
M_j	Number density of molecules of species j	(4.87)
\mathbf{M}	Magnetization (average magnetic dipole moment per unit volume)	(5.6)
$\mathbf{M}_{l,m}$	Vector spherical harmonic	(5.241) and (5.242)
\mathcal{M}_i	Mass of nucleus i in electron masses	(3.58)
$\mathcal{M}_j^{(1)}$	Number density of molecules of species j in ionization stage I	(4.87)
m	Mass of electron	
m_i	Mass of ion	
m_{amu}	Electron mass in atomic mass units	(3.40a)
N	Component of the total angular momentum at right angles to the internuclear axis. It represents the rotation of the nuclei alone	p. 61
N_b	Number of bound electrons per atom	(3.50)
N_o	Avogadro number	
N_e	Number of electrons per unit volume (see also ρ_e)	(3.24)
$N_e(r)$	Spherically averaged electron density as a function of r (distance from the nucleus)	(3.31a)
$N_e(v)$	Number of electrons with velocity v per unit volume	(5.96a) and (5.96b)
N_f	Number of free electrons per atom	(3.50)
N_I	Population of ionization stage I	(4.58)
N_i	Occupation number of level i	(2.2)
$N_i^{(n)}$	Occupation number of level i in a neutral atom	(3.47)
\bar{N}_i	Mean occupation number of level i	(3.44)
$\mathbf{N}_{l,m}$	Vector spherical harmonic	(5.241) and (5.243)
\mathcal{N}	Column number density	(7.140)
\mathcal{N}_{fi}	Occupation number of state i in ion I	(4.64)
\mathcal{N}_i^I	Number density of atoms of element i in ionization stage I	(4.85)
n	Complex refractive index	(5.19)
n_r	Real part of refractive index	(5.19)
n_i	Imaginary part of refractive index	(5.19)
$\hat{\mathbf{n}}$	Unit vector normal to a surface (Fig. 2.1)	(2.13)
n_i^*	Effective principal quantum number of level i	(3.53)
$n_\nu(\hat{\omega})$	Number of photons of frequency ν in direction $\hat{\omega}$	(5.54a) and (5.54b)
P_g	Gas pressure	(4.18)
P_i	Partial pressure of species i	(4.72a)
$P^{(c)}$	A priori probability for a configuration	(7.100)
$P_i(r)$	r times radial wave function of an electron in level i	(3.1)
$P_\nu(R)$	R times vibrational wave function of molecule	(3.76)
\mathbf{P}	Electric polarization (average electric dipole moment per unit volume)	(5.5)
\mathcal{P}_ν	Photoelectric edge distribution function	(6.3)
$\mathcal{P}_\nu(x)$	Symmetric band contour profile	(11.9)
$\mathcal{P}(v, \nu_o, \gamma, \Gamma)$	Pressure broadened Lorentz line profile	(7.191)

$p(\cos \theta)$	Phase function for angle between incident and scattered radiation	(2.9)
p_i	Ratio of occupation to statistical weight of state i	(2.2) and (3.41)
p_f	Probability that the final state is occupied	(6.1)
\mathbf{p}	Momentum of electron	(5.48)
\mathbf{p}_F	Electron Fermi momentum	(9.43)
$p_{lv}^{uv'}$	Band strength for diatomic molecular band	(5.125)
p_R	Radial momentum	(3.68) and (3.70)
p_θ	Angular momentum	(3.69)
Q	Partition function	(4.94)
Q_{el}	Electronic partition function	(4.94) and (4.95)
Q_I	Partition function of ionization state I	(4.59)
Q_{rot}	Rotational partition function	(4.94) and (4.97)
Q_{vib}	Vibrational partition function	(4.94) and (4.96)
Q	Number of nuclei in molecule	(3.56) and (3.58)
q	Number of bound levels	(3.2)
q	Number of equivalent electrons in a shell	(5.85b)
q	Autoionization line profile index	(5.160)
q	Number of electrons in a molecule	(3.56) and (3.57)
q_i, q_{ij}	Normal (generalized) coordinates	(5.112) and (5.116)
$q_{v'v}$	Franck–Condon factor	(5.120) and (5.122)
q_v	Space variable	(5.53), (5.54a), and (5.54b)
R	Internuclear separation in molecule	(3.60)
R_{ij}	Radiative rate between states i and j	(2.1)
R_e	Equilibrium separation of nuclei	(3.64b) and (5.109)
R_o	Universal gas constant	
R_{on}	Equilibrium separation of nuclei in molecule	(3.65)
R_{nl}	Radial part of wave function	(4.55a)
$\bar{R}_{v'v}$	R -centroid	(5.121)
R_∞	Rydberg constant	
R_i	Position coordinate of nucleus i	(3.56)
$ \mathcal{R} _{if}$	Electric dipole matrix element for a molecule from initial state i to final state f	(5.99)
r	Radius	
r_{ce}	Classical electron radius	(5.43)
r_D	Debye screening radius for ions	(4.61), (6.28), and (7.76)
r_{De}	Debye screening radius for electrons	(6.55)
r_i	Ion sphere radius for a fully ionized atom	(9.32)
r_n	Space coordinate normal to a surface	(2.16)
r_L	Lewis cutoff	(7.64)
r_o	Effective radius of atom	(4.2) and (4.4)
\mathbf{r}	Position coordinate of an electron	(5.14a), (5.14b), and (5.28)
$\mathbf{r}_e \equiv \mathbf{r}(R) _{lu}$	Electronic transition moment	(5.118)
S	Spin or multiplicity of state	(3.10)
S	Entropy	(4.19)
$S_{JK}^{J'K'}$	Hönl–London (H–L) factor	(5.101)
$S_{vJ}^{v'J'}$	Vibration-rotation line strength	(5.100)
$S_v^{v'}$	Vibration part of line strength	(5.105)

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Symbol	Description	Eq. No.
$S_J^{J'}$	Rotational factor of vibration-rotation line strength	(5.100)
S_{ij}	Transport coefficient	(9.12) and (9.13)
$\langle S \rangle$	Scattering integral	(9.55)
$S(\hat{\mathcal{L}})$	Relative line strength of a line in a multiplet	(5.67)
$S(\hat{\mathcal{M}})$	Relative multiplet strength of a multiplet in a transition array	(5.67)
$S(n, n')$	Electronic line strength	(5.75)
$[S_i]$	Concentration of species S_i	(4.80)
\mathbf{S}	Poynting vector	(5.35)
\hat{s}	Symmetry number of molecule	(4.97)
\hat{s}_i	Screening parameter of level i	(3.37)
T	Temperature	
T_D	Debye characteristic temperature	(9.40)
T_n	Electronic term value in molecule	(3.61b)
T_v	Period of molecular vibration	(3.68)
$T_v(\Delta r)$	Transmissivity at frequency ν	(2.61)
t	Time	
t_i	Energy scaling parameter	(3.48)
U	Internal energy per unit volume	(4.99)
\bar{U}	Energy density of electromagnetic field	(5.37)
u	Reduced photon energy	(2.7)
u_{rad}	Total radiation energy density	(2.48)
$u(m)$	Polarization bispinor for spin component m	(3.19)
$u_n(\mathbf{R}_i, \mathbf{r}_j)$	Electron eigenfunction for electronic state n in molecule	(3.57)
V_a	Volume of atom	(4.4)
$V_i(r)$	Potential energy for electron i	(3.14)
\mathcal{V}	Voigt line profile	(7.17)
$v_{n\nu J}(\mathbf{R}_i)$	Nuclear eigenfunction for electronic state n in molecule	(3.60)
\bar{v}	Mean thermal speed of electron	(9.5)
x_e	Vibrational constant	(3.62)
W_i	Relativistic energy of level i	(5.246)
$W(\mathbf{E})$	Holtzmark electric field distribution	(7.47)
$W_P(u)$	Planck weighting function	(2.59)
$W_R(u)$	Rosseland weighting function	(2.60)
w	Equivalent width of a line	(2.63)
w_{if}	Transition probability from initial state i to final state f	(5.45)
w_{jk}	Line coupling coefficient	(7.89)
$Y_n(x)$	Bessel function of the second kind (also known as Weber or Neumann functions) and order n	Appendix C (3.20)
$Y_{l,m}$	Spherical harmonic	(3.20)
$y_n(x)$	Spherical Neuman function of order n	(5.268c), Appendix C

Z	Atomic number	
Z'	Charge number of an ion	(3.53)
Z_i^*	Energy effective charge number of level i	(3.37)
Z_i^{**}	Wave function effective charge for level i	(5.86) and (5.89)
α	Chemical potential	(3.23)
α_e	Rotational-vibrational coupling constant	(3.64a)
α_M	Molar polarizability	(5.31b)
α_p	Polarizability	(5.29)
α_i	T-F scaling factor	(3.25)
α_o	Fine structure constant	(3.11)
$\tilde{\alpha}$	Line splitting per unit field strength	(7.70)
α_v	Absorptivity	(2.62)
α_v^1	Linear absorption coefficient wavenumber integrated over one line (SI units: m^{-2})	(7.140)
$\alpha_{vv'}$	Wavenumber integrated band absorption	(7.131)
$\alpha_{vJ}^{v'J'}$	Absorption coefficient wavenumber integrated over a single molecular line, corrected for stimulated emission	(7.1)
$\alpha_{vJ}^{v'}$	Absorption coefficient wavenumber integrated over rotational fine structure, corrected for stimulated emission	(7.2)
α	Dirac velocity matrix operator	(3.16)
$\bar{\alpha}_v$	Mean absorptivity wavenumber- integrated over a single line	(7.141)
β	Dirac matrix operator	(3.16)
β_i	Initial electron velocity divided by speed of light	(5.96b)
β_n	Spectroscopic constant	(3.65), (3.66a), and (3.66b)
Γ	Inelastic collision broadening half- maximum half-width (HMHW)	(7.22), (7.27), and (7.29)
Γ_C	Ion-ion Coulomb interaction parameter	(9.31)
$\tilde{\gamma}$	Lorentz contraction factor	(5.200)
γ_i	Activity coefficient of species i	(4.75)
$\gamma_i^{(n)}$	Natural radiative half-maximum half-width	(5.64)
γ_s	Ratio of the averaged squares of off-diagonal to diagonal components of scattering tensor	(5.222) and (5.223)
$\gamma^{(n)}$	Half-maximum half-width (HMHW) for medium with refractive index n	(5.26) and (5.41)
Δ_{ni}	Quantum defect	(3.53)
δ_l	Phase shift for l -wave function	(3.21a), (3.21b), and (3.55)
ϵ_o	Permittivity of vacuum	(5.5)
ϵ	Kinetic energy of free electron	(3.22)
ϵ_{dj}	Energy of formation of molecule j in an infinitely diluted gas	(4.104)
$\epsilon_v(\rho, T)$	Emissivity at frequency ν , density ρ , and temperature T . See also η_v	(2.32)
$\epsilon(\rho, T)$	Total emissivity at density ρ and temperature T	(2.38)

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Symbol	Description	Eq. No.
ϵ_c	Complex permittivity of medium	(5.18)
ϵ'_c	Real part of permittivity of medium	(5.18)
ϵ''_c	Imaginary part of permittivity of medium	(5.18)
ϵ^{LJ}	Depth of Lennard–Jones intermolecular potential	(8.45b)
$\tilde{\epsilon}_\nu$	Emission coefficient per unit volume at frequency ν	(2.3)
ϵ	Complex permittivity of a medium relative to that of vacuum (dielectric function)	(5.22a) and (5.23a)
ϵ	Electron energy	
ϵ_F	Electron energy at the Fermi surface	(9.45)
ξ_i	Spin–orbit radial interaction for level i	(3.11)
$\xi_i(x)$	Riccatti–Hankel function	(5.268b)
$\hat{\xi}_{\varphi}$	Spherical unit basis vector	(5.240)
$\eta = \alpha/(kT)$	Electron degeneracy parameter	(4.9)
η^*	Electron degeneracy parameter corrected for continuum lowering	(4.8) and (4.9)
η_i	Normal coordinate	(3.85)
η_ν	Emission efficiency at frequency ν ; see also $\epsilon_\nu(\rho, T)$	(5.257), (5.276a), and (5.276b)
Θ	Angle between incident and scattered radiation	Fig. 2.1
θ	Polar angle	Fig. 2.1
κ_e	Electron opacity	(9.2)
κ_ν	Mass extinction coefficient at frequency ν	(2.10b)
κ_P	Planck mean opacity	(2.57)
κ_R	Rosseland mean opacity	(2.58)
$\kappa_P(\Delta u)$	Planck group mean opacity	(2.64)
κ_t	Total (radiative and electron) Rosseland opacity	(9.4)
$\kappa_R(\Delta u)$	Rosseland group mean opacity	(2.65)
$\kappa_T(\Delta u)$	Transmission group mean opacity	(2.66)
Λ_e	Thermal conductivity by electrons	(9.1)
Λ_{ie}^s	Electron–ion scattering mean free path	(9.5)
Λ_L	Lorentz gas thermal conductivity	(9.11)
Λ_ν	Mean free path for extinction at frequency ν	(2.10a)
Λ_R	Transport equivalent of photon conductivity	(9.4a)
λ_C	Compton wavelength	(9.33)
μ_o	Permeability of vacuum	(5.6)
μ_m	Magnetic permeability	(5.9)
μ_r	Reduced nuclear mass	(3.64b)
$\tilde{\mu}_r$	Reduced nuclear mass in electron mass units	(3.76)
μ_a	Linear mean radiance weighted absorption coefficient	(2.52)
μ_f	Linear mean flux weighted absorption coefficient	(2.51)

μ_ν	Linear extinction coefficient at frequency ν	(2.5)
μ_ν^a	Linear absorption coefficient at frequency ν	(2.6a)
$\mu_\nu^{a'}$	Linear absorption coefficient at frequency ν corrected for stimulated emission	(2.5) and (2.6a)
μ_ν^s	Linear scattering coefficient at frequency ν	(2.5)
μ_ν^T	Linear transport extinction coefficient at frequency ν	(2.27)
μ_ν^{sT}	Linear transport scattering coefficient at frequency ν	(2.27)
$\mu_P(\rho, T)$	Planck mean absorption coefficient at density ρ and temperature T	(2.40)
$\mu_R(\rho, T)$	Rosseland mean extinction coefficient at density ρ and temperature T	(2.50)
$\tilde{\mu} = e r_e$	Atomic electric dipole moment	(5.28)
$\tilde{\mu} = e \mathcal{R}$	Molecular electric dipole moment	(5.99)
ν	Photon frequency	
$\tilde{\nu}$	Photon wavenumber	
ν_i	Stoichiometric coefficient of species i	(4.65)
ν_p	Electron plasma (or plasmon) frequency	(5.33)
ν_{pi}	Ion plasma frequency	(9.40)
ρ	Mass density	
ρ	Collision impact parameter	(7.24), (7.25), and (7.30)
ρ_e	Density of free electrons ($\rho_e = e N_e$)	(5.1)
ρ_2	Density operator	(7.96)
$\hat{\rho}$	Canonical (energy) density	(7.165)
$\hat{\rho}_f$	Density per unit volume of final states per energy interval	(5.45) and (5.52)
σ	Stefan–Boltzmann constant	(2.37a)
σ_e	Electric conductivity	(5.8) and (9.16)
σ_{ie}^s	Electron–ion scattering cross section	(9.8)
σ_p	Plasmon scattering cross section	(5.206a) and (5.206b)
σ_R	Rayleigh cross section for photon scattering	(5.44)
σ_T	Thomson cross section for photon scattering by free electrons	(5.43)
σ'	In general, photoabsorption cross section corrected for stimulated (induced) emission under LTE conditions $\sigma' = \sigma[1 - e^{-h\nu/(kT)}]$	
$\sigma_{if}^{(bb)}(\nu)$	Bound–bound cross section	(5.66a) and (5.66b)
$\sigma_{n_l^* n_l'^*}^{(bb,BS)}$	Burgess–Seaton b–b cross section	(5.73)
$\sigma_{n_l^* n_l'^*}^{(bf,BS)}$	Burgess–Seaton b–f cross section	(5.92)
$\sigma_i^{(bf,K)}(\nu)$	Kramers b–f semiclassical cross section	(5.86)
$\sigma_1^{(bf,St)}(\nu)$	Stobbe b–f cross section for hydrogenic ions	(5.87)
$\sigma^{(inel)}(\nu)$	Inelastic collision cross section	(7.34)
$\sigma^{(ff)}(\nu)$	Free–free cross section	(5.93a) and (5.93b)
$\sigma^{(ff,K)}(\nu)$	Kramers f–f semiclassical cross section	(5.96)
σ_ν	Total extinction cross section at frequency ν	(2.10a)

(continued)

(continued)

Symbol	Description	Eq. No.
σ_ν^s	Total scattering cross section at frequency ν	(2.9)
$\sigma_\nu^s(\Theta)$	Differential scattering cross section at frequency ν for phase angle Θ	(2.9)
σ_c^T	Transport scattering cross section including collective effects	(6.57)
σ_{ij}	Screening constant	(3.38)
$\sigma_{if}^{(bb)}$	Bound-bound cross section	(5.66a)
σ^{LJ}	Lennard-Jones molecular diameter or range	(8.23), (8.27), and (8.44)
$\hat{\sigma}$	Matrix element for oscillator strength of a transition array	(5.67)
τ	Lifetime of a state	(5.158)
τ	Relaxation time	(9.50)
τ_ν	Optical thickness for extinction at frequency ν	(2.31)
$\tau_\nu^{a'}$	Optical thickness for absorption at frequency ν	(2.31)
Φ	Azimuth angle	Fig. 2.1
$\Phi_{\kappa\eta\mu}$	Dirac bispinor	(3.18a) and (3.18b)
$\Phi_J(\theta, \phi, \chi)$	Rotator wave function for molecule	(3.60)
ϕ	Azimuthal angle	Fig. 2.1
$\phi(x)$	Thomas-Fermi function	(3.25b)
$\phi(\nu)$	Spectral function (see also $\phi(\nu)$)	(7.162) and (8.1)
χ_e	Complex electric susceptibility of medium	(5.10)
χ_m	Magnetic susceptibility of medium	(5.11)
χ_I	Ionization energy of ionization stage I	(4.58)
$\Psi(\mathbf{R}, \mathbf{r})$	Total wave function for molecule	(3.56)
Ψ_i	Dirac electron wave function	(3.16) and (3.19)
ψ	Electronic wave function	(5.50)
$\psi_I(x)$	Riccati-Bessel function	(5.268a)
$\psi_{n\nu}(R)$	R times radial part of vibrational wave function of electronic state n	(3.60)
$\Omega_{\kappa\mu}$	Spherical spinor	(3.18a) and (3.18b)
$\hat{\omega}$	Unit vector for direction	(2.3), Fig. 2.1
ω_e	Vibrational constant	(3.62) and (3.66a)
$\omega_e x_e$	Vibrational constant	(3.62) and (3.66b)
$\omega_e y_e$	Vibrational constant	(3.62)

Appendix B

Glossary and Abbreviations

Absorbance – The common logarithm (\log_{10}) of the reciprocal of the transmittance of a pure solvent = absorbancy = extinction.

Absortance – 1 minus the transmittance.

Absorption – When a photon is absorbed by an atom (molecules) the atom (molecule) changes to a higher state of energy.

Absorption coefficient – If the radiative flux through a material decreases with distance l in proportion to $e^{-\mu l}$, then μ is called the absorption coefficient (= extinction coefficient).

Absorption oscillator strength – The ratio of the observed equivalent width of an absorption line to the equivalent width predicted on the basis of the classical oscillator model.

Absorptivity – The constant a in Beer's law relation $A = a \cdot l \cdot c$, where A is the absorbance, l the path length, and c the concentration of solution. Also known as absorptive power. Formerly known as absorbency index, absorption coefficient, or extinction coefficient.

Activity – The ratio of the fugacity of a state to the fugacity of the reference state.

Activity coefficient – The ratio of the fugacity to the pressure of the real gas. A measure of the deviation of a real gas from an ideal gas.

Amagat density unit – Gas molecular density in g-moles/m³ at STP ($P = 1$ atmosphere and $T = 0^\circ\text{C}$). The amagat density is unique for every gas and varies with pressure and temperature. It automatically corrects the ideal gas law for a real gas. The amagat number = ρ/ρ_{STP} , where ρ is the measured density.

Amagat volume unit – Gas volume per 0.022 414 m³/g-mol, the volume one g-mol of an ideal gas at STP ($P = 1$ atmosphere and $T = 0^\circ\text{C}$).

AO – Atomic orbital.

APEX – Adjustable parameter exponential.

ATC theory – Anderson–Tsao–Curnutte theory for molecular line broadening and line coupling.

Auto-correlation function – Correlation of a function with itself. – See also correlation.

Avogadro constant – $N_o = 6.022\ 141\ 99 \times 10^{23}$ molecules/g-mol. – See also mole.

B–D – Bates–Damgard.

Beer–Lambert law (also known as Beer’s law) – $I = I_o \exp(\sigma Nl)$, where I_o is the radiance of the incident radiation, I the radiance of the transmitted radiation, σ is the total (extinction, i.e., absorption plus scattering) cross section for a single particle, N is the number density of particles, and l is the path length.

B–O – Born–Oppenheimer basis orbitals. Members of a complete set of independent orthogonal functions into which a wave function can be expanded.

CCA – Close-coupling approximation.

CCPPA – Coupled cluster polarization propagator approximation.

CEPA – Coupled electron-pair approximation.

C–G – Curtis – Godson approximation for molecular band absorption.

Chemical equilibrium – Implies that the composition does not change with time. If only one phase of matter exists (gas, liquid, or solid) then chemical equilibrium is homogeneous. If any two or all three phases are in equilibrium then chemical equilibrium is heterogeneous.

CHF – Coupled H–F.

CI – Configuration interaction.

CIA – Collision-induced absorption.

Clausius–Mosotti equation – (Mosotti also written: Mossotti) relates permittivity to polarizability. See also Lorentz–Lorenz formula.

CNDO – Complete neglect of differential overlap.

Collision parameter – $1/\sigma_{ic}^s = r_D/b_{90}$, the ratio of the Debye shielding length to the value of the impact parameter for which the deflection angle for scattering is 90° .

Collision logarithm – $\ln \sigma_{ic}^s$. See collision parameter.

Compton scattering – Inelastic (incoherent) high-energy scattering by a free electron.

Compton wavelength – $\lambda_C = h/(mc)$, where h is the Planck constant, c is the speed of light, and m is the mass of the particle. For an electron $\lambda_C = 2.426\ 3 \times 10^{-12}$ m.

Conductive opacity – See electron opacity.

Correlation of two functions $f_1(t)$ and $f_2(t)$ – $C(t) = \int_{-\infty}^{\infty} f_1^*(t) f_2(t + \tau) d\tau$, where the asterisk denotes the complex conjugate.

CPA – Coupled pair approximation.

CSF – Configuration state function.

Curve of growth (Wachstumskurve) – Relates how the equivalent width of a spectral absorption (or emission) line increases with the number of atoms or molecules in their initial state producing the line as a function of optical depth or path length for a given optical depth. – See also equivalent width.

Dalton's law – At constant temperature the total pressure exerted by a mixture of gases in a definite volume is equal to the sum of the individual pressures that each gas would exert if it occupied the same total volume alone. Discovered by Cavendish (1781), announced by Dalton (1810).

DCA – Detailed configuration accounting.

DCAETS – Detailed configuration accounting with explicit term splitting.

De Broglie wavelength – $\lambda_e = h/p$, where h is the Planck constant and $p = mv$ is the momentum of the particle.

DDA – Discrete dipole approximation.

Detailed balance – The principle of detailed balance refers to the exact balance between a particular process and its inverse process in a system that is in thermodynamic equilibrium. When results of this principle are applied to a system that is not in strict thermodynamic equilibrium, then it is important that an appropriately defined thermodynamic quasi- or restricted equilibrium is applied (e.g., certain energy modes that equilibrate quickly as opposed to other energy modes that equilibrate much more slowly) so that the appropriate temperature can be defined.

DFT – Density functional theory.

Diffraction – Modification of radiation as it passes edges of opaque bodies.

Dipole polarizability – See polarizability.

Dispersion – Change of wavelength of radiation passing through a medium.

Disproportionation – Simultaneous oxidation and reduction of a substance reacting with itself to form two dissimilar molecules.

DTA – Detailed term accounting. Also referred to as DCAETS.

Effective band width – $\Delta\tilde{\nu}_{\nu-\nu'}$ corresponds to the wave numbers bounding the spectral region from which appreciable radiant energy is emitted or absorbed.

Einstein coefficient A – Spontaneous emission coefficient.

Einstein coefficient B – Coefficient for induced (stimulated) emission defined here in terms of radiation energy density per unit frequency interval.

Electron conduction opacity – See electron opacity.

Electron opacity – The measurable ability of a substance to obstruct the flow of thermal energy carried by electrons.

Emission (spontaneous) – An atom (molecule) in an excited state (energy state higher than the ground state) emits a photon without fixed phase relationship to other photons emitted by other atoms (molecules) in identically excited states.

Emission (induced) – An atom (molecule) in an excited state emits a photon of frequency ν under the action of an electromagnetic field at the same frequency. The incident and emitted photons have the same phase and the number of photons in the radiation field at frequency ν has increased by 1.

Emissivity – The ratio of radiation emitted by a surface (exitance) at a given frequency and temperature to the radiation emitted by a perfect blackbody radiator at the same frequency and temperature.

Energy density – See radiation energy density.

Enthalpy – The thermodynamic potential $H = U + PV$, where U is the internal energy, P is the pressure of the system, and V is the volume.

Entropy – A measure of the amount of energy in a physical system that *cannot* be used to do work. It also is a measure of disorder present in a system.

EOS – Equation of state.

Equilibrium – The state of a system in which the composition and properties undergo no observable changes.

Equivalent line width – The width of an imaginary line with a rectangular profile that is completely absorbing (or emitting Planckian radiance) equivalent to a real line (or molecular band) that absorbs (or emits) the same number of photons.

Exciton – A quantum of electronic excitation consisting of an electron–hole pair.

Exitance – The rate at which an extended source radiates energy *in all directions* per unit area. Units are W m^{-2} .

Fermi’s “Golden Rule No. 2” – Originally formulated by Dirac (1927a). In perturbation theory the time rate for a transition of a system from initial state i to a final state f .

F–C – Franck–Condon.

Flux (radiative flux) – The rate at which energy is radiated from a source. Units are W . See also normal flux density.

Flux per unit frequency (radiative flux per unit frequency) – The rate at which energy is radiated from a source at a given frequency. Units are W Hz^{-1} . See also normal flux density per unit frequency.

FOCI – First order configuration interaction.

Free energy – See Gibbs free energy and Helmholtz free energy.

Fugacity – The measure for the tendency of a gas to escape or expand. It is the pressure value needed at a given temperature to make the properties of a non-ideal gas satisfy the equation for an ideal gas, i.e., $f_i = \gamma_i P_i$, where γ_i is the fugacity coefficient and P_i is the partial pressure for component i of the gas. For an ideal gas, $\gamma_i = 1$.

Gaunt factor – Ratio of quantum mechanical to classical (Kramers or hydrogenic) photo cross section.

Gibbs free energy – A measure of the available work that can be extracted from some process operating at constant pressure. $G = H - TS$, where H is the enthalpy, T the temperature, and S the entropy.

GLFE – Global-local finite element.

HAM – Hydrogenic atoms in molecules.

Helmholtz free energy – $F = \bar{E} - TS$, where \bar{E} is the mean value of the total energy at temperature T and density ρ and S is the entropy.

Heterogeneous system – A system of two or more regions of homogeneity.

H-F – Hartree–Fock.

H-F-S – Hartree–Fock–Slater.

H-L – Hönl–London.

HMHW – Half maximum half width: Half width at half of the peak of a spectral line.

HNC – Hypernetted chain theory.

Homogeneous system – A system whose properties are uniform throughout, or vary in a continuous manner.

HX – Hartree-plus-statistical-exchange.

IHNC – Improved hypernetted chain (HNC) theory.

INDO – Intermediate neglect of differential overlap.

Intensity (radiative intensity) – The radiant flux per unit solid angle emitted by a radiating source in a specific direction, $\hat{\omega}$. Units are W sr^{-1} .

IR – Infra red.

Irradiance – Radiative power per unit area incident on a surface. Units are W m^{-2} . If the angle of incidence is normal to the surface, then irradiance equals normal flux density.

J-W-K-B – Jeffrey–Wenzel–Kramers–Brillouin.

Kinetic equilibrium – Processes in which chemical reactions lead to products that are separated from true thermodynamic (including chemical) equilibrium by a potential barrier and the temperature is too low to overcome the barrier except at extremely low rates.

Kirchhoff's law – Expresses the general principle of detailed balance as applied to the emission and absorption of radiation.

Klein-Nishina scattering – Relativistic (inelastic, incoherent) photon scattering by a free electron at rest.

Law of mass action – The rate of an elementary reaction (defined by reduction of reactant or formation of product) is proportional to the concentration of each individual species involved in the elementary reaction.

LCAO – Linear combinations of atomic orbitals.

Line strength – Absorption coefficient integrated over frequency, wave number, or wavelength.

Lorentz–Lorenz formula – Relates the refractive index of a dilute gas to its density (temperature and pressure) and the molar refraction. See Clausius–Mosotti equation.

Lorentz gas – A plasma in which electrons only interact with ions. Electron–electron interaction is ignored.

Lorenz number – $L = 4\pi\epsilon_0(\pi k)^2/(3e^2) = 2.718\,215 \times 10^{-18} \text{ J m}^{-1}\text{K}^{-2} = 2.44 \times 10^{-8} \text{ V}^2\text{K}^{-2}$. See Wiedemann–Franz law.

LTE – Local thermodynamic equilibrium.

LTHF – Linearized time-dependent Hartree–Fock.

Magnon – Quantized magnetic spin wave.

MBPT – Many-body perturbation theory.

MCA – Multi-configuration approximation.

MECI – Mono-excited configuration interaction.

MHD – Magnetohydrodynamic.

MIUTS – Mean ion with unfolded term splitting.

MLG – Multiple line group model for molecular band absorption.

MO – Molecular orbital.

Molal properties – Properties for which values are expressed per mole of the compound whose chemical formula is written.

Molar concentration – Number density.

Molar Gibbs free energy – The partial derivative of the Gibbs free energy with respect to \hat{n}_i , the number of moles of a constituent in a system. Molar Gibbs free energy is also called the partial molar Gibbs free energy or the molar chemical potential (or briefly, the chemical potential).

Molar polarizability – $\alpha_M = \frac{\epsilon-1}{\epsilon+2} \frac{10^{-3}M}{\rho}$, where $\epsilon = \epsilon_e/\epsilon_0$.

Molar refraction – (in SI units) $R = (10^{-3}M/\rho) \cdot (n^2 - 1)/(n^2 + 2)$, where M is the gram-molecular weight, ρ the density, and n the refractive index.

Molar volume – The volume occupied by one mole of ideal gas at standard temperature and pressure (STP). Its value is $2.241\,4 \times 10^4 \text{ m}^3/\text{g-mol}$.

Mole (also gram-mole) – $1 \text{ g-mol} = 1 \text{ gram molecular weight} (= 10^{-3} \text{ kg molecular weight})$.

Mole fraction – The number of moles of a component of a solution divided by the total number of moles of all components.

MRD – Multireference double-excitation.

NDDO – Neglect of diatomic differential overlap.

Neutrino opacity – The measurable ability of a substance to obstruct the flow of energy carried by neutrino pairs.

Normal flux density – The rate of radiative energy flow per unit area normal to the surface of a sphere enclosing the radiation source. Units are W/m^2 .

Normal flux density per unit frequency – The rate of radiative energy flow at a given frequency per unit area normal to the surface of a sphere enclosing the radiation source. Units are $W/(m^2 Hz)$.

NLTE – Non-local thermodynamic equilibrium.

Occupation number – The number of electrons in a level of an atom or molecule.

ODF – Opacity distribution function.

Opacity – The measurable ability of a substance to obstruct by absorption and scattering the transmission of radiant energy. Opacity thus is the degree of non-transparency.

Optical depth – The dimensionless line integral of the absorption, scattering, or extinction coefficient along any path in a radiation transmitting medium.

Parity – Property change of a wave function when its variables are replaced by the negative of their coordinates.

Partition function – The primary quantity in statistical mechanics from which all thermodynamic quantities can be calculated.

Phonon – Quantized lattice wave.

PIA – Pressure-induced absorption. The same as collision-induced absorption (CIA).

Plasma frequency – The plasma frequency is the natural frequency of oscillation of electrons in a plasma displaced relative to the ion background.

Plasmon – Quantum of charge-density oscillations in a plasma; a charge-density wave in an electrically neutral collection of charges in which some of the charges are free to move in response to their Coulomb interaction.

Polariton – Quantum of coupled electromagnetic wave and another excitation (e.g., a phonon, plasmon, etc.).

Polarizability (dipole) – A microscopic electric property of dielectric media that relates the macroscopically observable electric susceptibility to the properties of the atom or molecule of the medium. It is a second rank tensor that in lowest order is a measure of the response (distortion) of the field strength of an electron cloud to an external field. It measures the average local polarization. Hyperpolarizability is a fourth-rank tensor in the study of nonlinear effects in atoms. It should not be confused with higher-order multipoles of the atomic charge distribution.

Polarization (atomic) – Induced displacement of an electron cloud toward the stronger atom in a molecule under the influence of an applied electric field.

Polarization (electronic) – Induced displacement of electrons relative to their nuclei under the influence of an applied electric field.

Polarization (interfacial) – Induced but limited migration of charge carriers (space-charge polarization).

Polarization (orientation) – See polarization (permanent).

Polarization (permanent) – Asymmetric charge distribution between unlike atomic partners of a molecule.

Poynting vector – $\mathbf{S} = \mathbf{E} \times \mathbf{H}$, where \mathbf{E} is the electric field and \mathbf{H} is the magnetic field. See also time-averaged Poynting vector.

QMD – Quantum Molecular Dynamics.

Radiance – The rate at which radiative energy passes through an area or is emitted from an area within a specific solid angle in a specific direction $\hat{\omega}$. Units are $\text{W sr}^{-1} \text{m}^{-2}$. (In astronomy, astrophysics, and heat transfer, radiance is also incorrectly called specific intensity, or just intensity. See definition of intensity).

Radiation (or radiative) energy density – Total radiation energy in a field per unit volume and per unit frequency. It is evaluated by taking the density of states in a volume, $\rho(E)$ (our $\hat{\rho}_f$), multiplying by the energy per state, $Nh\nu$, dividing by the volume, V , and accounting for the two polarization states by multiplying by 2.

Radiative flux – See flux.

Raoult's law – The vapor pressure of a solvent in an ideal solution equals the mole fraction of the solvent times the vapor pressure of the pure solvent.

Reflectance – The ratio of reflected flux to incident flux from a dielectric.

Reflectivity – The fraction of incident radiation that is reflected from a body.

Refraction – Change of speed of light, and hence of the wavelength and direction of radiation, associated with a change of the refractive index.

Resonance fluorescence – The two-photon process between the ground state and the first excited state.

R–H–F – Restricted Hartree–Fock.

R–K – Rydberg–Klein.

R–K–R – Rydberg–Klein–Rees.

RPA – Random phase approximation.

RPAE – Random phase approximation with exchange.

Saha ionization equation – Relates the fraction of ionized atoms in a plasma as a function of the temperature, T , and electron density of the plasma.

SCA – Statistical Configuration Accounting (see also UTA and STA).

SCEP – Self-consistent electron-pair.

SCF – Self-consistent-field.

SD – Calculation at a singles – doubles level.

SDQ – Singles – doubles quadrupole.

SHO – Simple harmonic oscillator.

SOPPA – Second order polarization propagator approximation.

SOS – Statistical opacity sampling.

Spectral irradiance – Radiative power per unit area and per unit frequency (or unit wave number) incident on a surface.

Spectral radiance – The rate at which radiative energy at a given frequency (or wavenumber) passes through an area or is emitted from an area within a specific solid angle in a specific direction $\hat{\omega}$. Units are $\text{W sr}^{-1} \text{m}^{-2} \text{Hz}^{-1}$ (or $\text{W sr}^{-1} \text{m}^{-3}$). (In astronomy, astrophysics, and heat transfer, radiance is incorrectly called specific spectral intensity, spectral intensity, or just intensity. See definition of intensity).

STA – Super transition array. A STA is the total transition array of a specific single-electron transition, including all possible contributing configurations, between a pair of superconfigurations.

State of a system – A state whose properties and composition is known.

Steady state – The *sum* of all processes producing a particular state is balanced by the *sum* of all processes destroying it, but the producing processes are not inverse of the destruction processes.

STO – Slater-type orbital.

Stoichiometric coefficient – Represents the quantitative degree of a chemical species participating in a reaction.

Strong coupling regime – The ratio of Coulomb energy of ions in a dense plasma relative to the thermal energy of the plasma, kT , is larger than one.

Superconfiguration – The collection of all ordinary configurations obtained by distributing N_i electrons to the ordinary shells of a supershell in all possible ways allowed by the Pauli exclusion principle.

TDA – Tamm–Dancoff approximation.

TDHF – Time-dependent H–F.

T–F – Thomas–Fermi.

T–F–A – Thomas–Fermi–Amaldi.

T–F–D – Thomas–Fermi–Dirac.

Thermodynamic equilibrium – Implies three different simultaneous equilibria: (1) Thermal equilibrium, for which the temperature must be the same throughout the system. (2) Mechanical equilibrium, which means no macroscopic movements within the system or of the system with respect to its surroundings. (3) Chemical equilibrium, which implies that the composition does not change with time. In thermodynamic equilibrium each process is exactly balanced by its inverse process. (See also detailed balance).

Thermodynamic process – A process in which changes in a system take place when subjected to thermodynamic examination.

Thermodynamic system – Part of the universe that is subjected to thermodynamic scrutiny.

Thomson scattering – Elastic (coherent) low-energy photon scattering by a free electron.

Time-averaged Poynting vector for harmonic fields – Radiation energy flux density.

Transition array – The totality of lines resulting from transitions between two configurations.

Transition cluster – The resulting spectrum corresponding to a one-electron transition from the lower level $(n, l, j)_l$ to the upper level $(n, l, j)_u$. It may contain many transition arrays.

Transmission coefficient – The ratio of the transmitted flux of radiation to the incident flux for a substance of unit thickness.

Transmissivity – The fraction of incident radiation transmitted through a body.

Transmittance – The radiant power transmitted by a body divided by the total radiant power incident upon the body (also known as transmission).

UTA – Unresolved transition array.

UV – Ultra violet.

VCI – Valence configuration interaction.

Vibrational angular momentum – A feature of linear polyatomic molecules. For CO_2 , for example, the vibrational angular momentum quantum number l measures the angular rotation in units of $h/(2\pi)$, which is associated with the degenerate (bending) ν_2 vibration. l is usually written as a superscript to the corresponding vibrational quantum number, e.g., $\nu_1\nu_2^l\nu_3$. (See [Herzberg 1945](#).) Rotational energies and selection rules have l -dependencies.

V-R – Vibration–rotation.

Wiedemann–Franz law – Also known as the Wiedemann–Franz–Lorenz law, relates electric and thermal conductivity via the Lorenz number when a single, unique relaxation time can be calculated for non-relativistic electrons. See Lorenz number.

Wigner D function – Matrix representation of the finite rotation operator using spherical harmonics.

Appendix C

Some Mathematical Functions

C.1 Bessel Functions

The functions $J_p(z)$ and $Y_p(z)$ are the solutions of the first and second kind, respectively, to the Bessel differential equation

$$z^2 \frac{d^2 w}{dz^2} + z \frac{dw}{dz} + (z^2 - p^2) w = 0,$$

where $z = x + iy$ and p is a real number.

C.1.1 Bessel Functions of the First, Second, and Third Kind

Bessel Functions of the First Kind

$$\begin{aligned}
 J_p(z) &= \sum_{k=0}^{\infty} \frac{(-1)^k}{k! \Gamma(p+k+1)} \left(\frac{z}{2}\right)^{p+2k}, \\
 &\quad (\text{parameter } p \text{ is a real number, } 0 < z < \infty). \\
 &= \frac{1}{\pi} \int_0^\pi \cos(z \sin \phi - p\phi) d\phi - \frac{\sin(p\pi)}{\pi} \int_0^\infty e^{-z \sinh t - pt} dt, \\
 &\quad (|\arg z| < \frac{\pi}{2}).
 \end{aligned}$$

$$J_n(z) = \sum_{k=0}^{\infty} \frac{(-1)^k}{k!(n+k)!} \left(\frac{z}{2}\right)^{n+2k},$$

$$= \frac{1}{\pi} \int_0^{\pi} \cos(z \sin \phi - n\phi) d\phi, \quad (n = 0, 1, 2, \dots).$$

$$J_{-n}(z) = (-1)^n J_n(z), \quad (n = 1, 2, 3, \dots).$$

$$\frac{\partial J_p(z)}{\partial z} = \frac{1}{2} [J_{p-1}(z) - J_{p+1}(z)].$$

In particular:

$$J_0(z) = 1 - \frac{\left(\frac{z}{2}\right)^2}{(1!)^2} + \frac{\left(\frac{z}{2}\right)^4}{(2!)^2} - \frac{\left(\frac{z}{2}\right)^6}{(3!)^2} + \dots,$$

$$J_1(z) = \frac{z}{1!} - \frac{\left(\frac{z}{2}\right)^3}{1!2!} + \frac{\left(\frac{z}{2}\right)^5}{2!3!} - \frac{\left(\frac{z}{2}\right)^7}{3!4!} + \dots,$$

$$J'_0(z) \equiv \frac{\partial J_0(z)}{\partial z} = -J_1(z).$$

Bessel Functions of the Second Kind (Also Called Weber or Neumann Functions)

$$Y_p(z) = \frac{J_p(z) \cos(p\pi) - J_{-p}(z)}{\sin(p\pi)},$$

(parameter p is a real number, $p \neq 0, 1, 2, \dots$),

$$Y_p(z) = \frac{1}{\pi} \int_0^{\pi} \sin(z \sin \phi - p\phi) d\phi - \frac{1}{\pi} \int_0^{\infty} [e^{pt} + e^{-pt} \cos(p\pi)] e^{-z \sinh t} dt,$$

$$(|\arg z| < \frac{\pi}{2}).$$

$$Y_n(z) = \frac{2}{\pi} \ln\left(\frac{z}{2}\right) J_n(z) - \frac{1}{\pi} \sum_{k=0}^{n-1} \frac{(n-k-1)!}{k!} \left(\frac{z}{2}\right)^{2k-n}$$

$$- \frac{1}{\pi} \sum_{k=0}^{\infty} (\psi(k+1) + \psi(k+n+1)) \frac{(-1)^k}{k!(n+k)!} \left(\frac{z}{2}\right)^{2k+n},$$

$$= \frac{1}{\pi} \int_0^\pi \sin(z \sin t - nt) dt - \frac{1}{\pi} \int_0^\infty [e^{nt} + (-1)^n e^{-nt}] e^{-z \sinh t} dt,$$

$$(|\arg z| < \frac{\pi}{2}, \quad n = 0, 1, 2, \dots),$$

where $\psi(m)$ is the digamma function: $\psi(1) = -\gamma$, $\psi(m) = -\gamma + \sum_{j=1}^{m-1} (j^{-1})$, and $\gamma = 0.577\,215\,665 \dots$ is Euler's constant.

$$Y_{-n}(z) = (-1)^n Y_n(z), \quad (n = 1, 2, 3, \dots).$$

$$\frac{\partial Y_p(z)}{\partial z} = \frac{1}{2} [Y_{p-1}(z) - Y_{p+1}(z)].$$

In particular:

$$Y_0(z) = \frac{2}{\pi} \left\{ \left[\ln \left(\frac{z}{2} \right) + \gamma \right] J_0(z) + \frac{\left(\frac{z}{2} \right)^2}{(1!)^2} - \frac{\left(1 + \frac{1}{2} \right) \left(\frac{z}{2} \right)^4}{(2!)^2} + \frac{\left(1 + \frac{1}{2} + \frac{1}{3} \right) \left(\frac{z}{2} \right)^6}{(3!)^2} - \dots \right\},$$

$$Y'_0(z) \equiv \frac{\partial Y_0(z)}{\partial z} = -Y_1(z).$$

Bessel Functions of the Third Kind (Also Called Hankel Functions)

$$H_p^{(1)}(z) = J_p(z) + i Y_p(z),$$

$$H_p^{(2)}(z) = J_p(z) - i Y_p(z),$$

$$H_{-p}^{(1)}(z) = e^{p\pi i} H_p^{(1)}(z),$$

$$H_{-p}^{(2)}(z) = e^{-p\pi i} H_p^{(2)}(z), \quad (\text{parameter } p \text{ is a real number}).$$

$$H_n(z) = \sum_{k=0}^\infty \frac{(-1)^k}{k!(n+k)!} \left(\frac{z}{2} \right)^{n+2k}, \quad (n = 0, 1, 2, \dots).$$

$$= \frac{1}{\pi} \int_0^\infty \cos(z \sin \phi - n\phi) d\phi$$

$$= \frac{1}{2\pi} \int_{-\pi}^\pi e^{i(z \sin \phi - n\phi)} d\phi.$$

C.1.2 Modified Bessel Functions (Also Called Basset Functions) of the First and Second Kind

The functions $I_p(z)$ and $K_p(z)$ are solutions of the first and second kind, respectively, to the modified Bessel differential equation

$$z^2 \frac{d^2 w}{dz^2} + z \frac{dw}{dz} - (z^2 + p^2) w = 0,$$

where $z = x + iy$ and p is a real number.

Modified Bessel Functions of the First Kind

$$I_p(z) = \sum_{k=0}^{\infty} \frac{1}{k! \Gamma(p+k+1)} \left(\frac{z}{2}\right)^{p+2k}, \quad (\text{parameter } p \text{ is a real number}).$$

$$I_n(z) = \sum_{k=0}^{\infty} \frac{1}{k!(n+k)!} \left(\frac{z}{2}\right)^{n+2k}, \quad (n = 0, 1, 2, \dots).$$

$$I_n(z) = e^{-in\pi/2} J_n(ze^{i\pi/2}), \quad (-\pi < \arg z \leq \frac{\pi}{2}),$$

$$= e^{i3n\pi/2} J_n(ze^{-i3\pi/2}), \quad (\frac{\pi}{2} < \arg z \leq \pi),$$

$$= \frac{1}{\pi} \int_0^{\pi} e^{z \cos \phi} \cos(n\phi) d\phi.$$

$$I_{-n}(z) = I_n(z), \quad (n = 1, 2, 3, \dots).$$

In particular:

$$I_0(z) = 1 + \frac{(\frac{z}{2})^2}{(1!)^2} + \frac{(\frac{z}{2})^4}{(2!)^2} + \frac{(\frac{z}{2})^6}{(3!)^2} + \dots,$$

$$I_1(z) = \frac{z}{2} + \frac{(\frac{z}{2})^3}{1!2!} + \frac{(\frac{z}{2})^5}{2!3!} + \frac{(\frac{z}{2})^7}{3!4!} + \dots,$$

$$I_0'(z) \equiv \frac{\partial I_0(z)}{\partial z} = I_1(z).$$

Modified Bessel Functions of the Second Kind

$$K_p(z) = \frac{\pi[I_{-p}(z) - I_p(z)]}{2 \sin(p\pi)}, \quad (\text{parameter } p \text{ is a real number}).$$

$$K_p(z) = \frac{\pi}{2 \sin(p\pi)} \left[\sum_{k=0}^{\infty} \frac{1}{k!(n+k)!} \left(\frac{z}{2}\right)^{-n+2k} - \sum_{k=0}^{\infty} \frac{1}{k!(n+k)!} \left(\frac{z}{2}\right)^{n+2k} \right],$$

$$\begin{aligned} K_n(z) &= \frac{i\pi}{2} e^{in\pi/2} H_n^{(1)}(ze^{i\pi/2}), \quad (-\pi < \arg z \leq \frac{\pi}{2}), \\ &= -\frac{i\pi}{2} e^{-in\pi/2} H_n^{(2)}(ze^{-i\pi/2}), \quad (-\frac{\pi}{2} \arg z \leq \pi). \end{aligned}$$

$$K_{-n}(z) = K_n(z), \quad (n = 1, 2, 3, \dots).$$

$$\frac{\partial K_p(z)}{\partial z} = \frac{1}{2} [K_{p-1}(z) + K_{p+1}(z)].$$

In particular:

$$K_0(z) = -\left[\ln\left(\frac{z}{2}\right) + \gamma\right] I_0(z) + \frac{(\frac{z}{2})^2}{(1!)^2} + \frac{(1 + \frac{1}{2})(\frac{z}{2})^4}{(2!)^2} + \frac{(1 + \frac{1}{2} + \frac{1}{3})(\frac{z}{2})^6}{(3!)^2} + \dots$$

$$\begin{aligned} K_1(z) &= \frac{1}{z} + \frac{z}{4} \left[2\gamma - 1 + \frac{1}{8} \left(2\gamma - \frac{5}{2} \right) z^2 + \frac{1}{192} \left(2\gamma - \frac{10}{3} \right) z^4 + \dots \right] \\ &\quad + \frac{z}{2} \ln \frac{z}{2} \left(1 + \frac{z^2}{8} + \frac{z^4}{192} + \dots \right) \end{aligned}$$

$$K'_0(z) \equiv \frac{\partial K_0(z)}{\partial z} = -K_1(z).$$

$$\lim_{z \rightarrow 0} z K_1(z) \rightarrow 1,$$

$$\lim_{z \rightarrow \infty} K_1(z) \rightarrow \left(\frac{\pi}{2z}\right)^{1/2}.$$

C.1.3 Spherical Bessel Functions

The functions $j_n(z)$ and $y_n(z)$ are the solutions of the first and second kind, respectively, to the Bessel differential equation

$$z^2 \frac{d^2 w}{dz^2} + 2z \frac{dw}{dz} + [z^2 - n(n+1)]w = 0,$$

where $z = x + iy$ is a real number and n are integers.

Spherical Bessel Function of the First Kind

$$j_n(z) = \sqrt{\frac{\pi}{2z}} J_{n+1/2}(z)$$

In particular:

$$j_0(\rho) = \frac{\sin(\rho)}{\rho},$$

$$j_1(\rho) = \frac{\sin(\rho)}{\rho^2} - \frac{\cos(\rho)}{\rho},$$

$$j_2(\rho) = \left(\frac{3}{\rho^2} - 1\right) \frac{\sin \rho}{\rho} - \frac{3 \cos \rho}{\rho^2},$$

$$j_3(\rho) = \left(\frac{15}{\rho^3} - \frac{6}{\rho}\right) \frac{\sin \rho}{\rho} - \left(\frac{15}{\rho^2} - 1\right) \frac{\cos \rho}{\rho}.$$

Spherical Bessel Function of the Second Kind

$$y_n(z) = \sqrt{\frac{\pi}{2z}} Y_{n+1/2}(z)$$

In particular:

$$y_0(\rho) = -\frac{\cos(\rho)}{\rho},$$

$$y_1(\rho) = -\frac{\cos(\rho)}{\rho^2} - \frac{\sin(\rho)}{\rho},$$

$$y_2(\rho) = \left(-\frac{3}{\rho^2} + 1\right) \frac{\cos \rho}{\rho} - \frac{3 \sin \rho}{\rho^2},$$

$$y_3(\rho) = \left(-\frac{15}{\rho^3} + \frac{6}{\rho}\right) \frac{\cos \rho}{\rho} - \left(\frac{15}{\rho^2} - 1\right) \frac{\sin \rho}{\rho}.$$

Spherical Bessel Functions of the Third Kind (Spherical Hankel Functions)

$$h_n^{(1)}(\rho) = j_n(\rho) + iy_n(\rho),$$

$$h_n^{(2)}(\rho) = j_n(\rho) - iy_n(\rho).$$

In particular:

$$h_0^{(1)}(\rho) = \frac{e^{i\rho}}{i\rho},$$

$$h_1^{(1)}(\rho) = -\frac{e^{i\rho}}{\rho} \left(1 + \frac{i}{\rho}\right),$$

$$h_2^{(1)}(\rho) = \frac{ie^{i\rho}}{\rho} \left(1 + \frac{3i}{\rho} - \frac{3}{\rho^2}\right).$$

C.2 Binomial Function (Hypergeometric Distribution)

$$f(x) = \frac{\binom{M}{x} \binom{N-M}{n-x}}{\binom{N}{n}},$$

where

$$\binom{M}{x} \equiv \frac{M!}{x!(M-x)!}, \quad (0 \leq x \leq M).$$

Here $f(x)$ is the probability of obtaining exactly x elements of one kind and $n - x$ elements of another, if n elements are chosen at random without replacement from a finite population containing N elements of which M are of the first kind and $N - M$ are of the second kind.

C.2.1 Binomial Distribution

When there are exactly two mutually exclusive outcomes of a trial: “success” and “failure,” the binomial distribution gives the probability of observing x successes in N trials, with the probability of success on a single trial denoted by p . The binomial probability function is

$$P(x, p, N) = \binom{N}{x} p^x (1-p)^{(N-x)}, \quad (x = 0, 1, 2, \dots, N).$$

The binomial cumulative probability function is

$$F(x, p, N) = \sum_{i=0}^x \binom{N}{i} p^i (1-p)^{(N-i)}.$$

C.3 Fermi Integrals

C.3.1 Non-relativistic Fermi Integrals

$$\begin{aligned} F_k(\eta) &\equiv \frac{1}{\Gamma(k+1)} \int_0^\infty \frac{x^k}{e^{x-b\eta} + 1} dx, \\ F'_k(\eta) &\equiv \frac{\partial}{\partial \eta} I_k(\eta) \\ &= \frac{1}{\Gamma(k+1)} \int_0^\infty x^k \frac{\partial}{\partial \eta} \left(\frac{1}{e^{x-\eta} + 1} \right) dx \\ &= -\frac{1}{\Gamma(k+1)} \int_0^\infty x^k \frac{\partial}{\partial x} \left(\frac{1}{e^{x-\eta} + 1} \right) dx \\ &= -x^k \frac{1}{e^{x-\eta} + 1} \Big|_0^\infty + k \int_0^\infty \frac{x^{k-1}}{e^{x-\eta} + 1} dx \\ &= k F_{k-1}(\eta). \end{aligned}$$

In particular:

$$\frac{\partial}{\partial \eta} F_{1/2}(\eta) = F_{-1/2}(\eta),$$

For additional information and tabulations see [Rhodes \(1950\)](#), [Dingle \(1957\)](#), and [Wyller \(1973\)](#).

Approximations

$$F_{1/2}(\eta) = \frac{2}{3}\eta^{3/2} \left(1 + \frac{1.23}{\eta^2} \right), \quad (\eta > 30),$$

$$F_{1/2}(\eta) = \frac{2}{3}\eta^{3/2} \left(1 + \frac{1.23}{\eta^2} + \frac{1.25}{\eta^4} \right), \quad (4 \leq \eta \leq 30),$$

$$F_{1/2}(\eta) = (0.008\eta^3 + 0.187\eta^2 + 0.525\eta + 0.678) \\ + (-0.012\eta^2 + 0.019\eta - 0.009)|\eta|, \quad (-2 \leq \eta < 4),$$

$$F_{1/2}(\eta) = 0.8862e^\eta - 0.3133e^{2\eta} + 0.569e^{3\eta}, \quad (-28 < \eta < -2),$$

$$F_{1/2}(\eta) = e^{(\eta-0.12078223)}, \quad (\eta \leq -28),$$

$$F_{3/2}(\eta) = \frac{3}{4}\pi^{1/2}e^\eta(1 - 0.167764e^\eta + 0.064150027e^{2\eta} - 0.03125e^{3\eta} \\ + 0.017885431e^{4\eta} - 0.011340229e^{5\eta} + 0.00771356158e^{6\eta}), \\ (\eta \leq -2),$$

$$F_{3/2}(\eta) = 0.1758009896 + \frac{3}{2}(0.6513262112 + 0.678091\eta + 0.268098335\eta^2 \\ + 0.05636582666\eta^3 + 0.00469520575\eta^4 - 0.00047150892\eta^5 \\ - 0.0001066017995\eta^6), \quad (-2 < \eta \leq 0),$$

$$F_{3/2}(\eta) = 1.152790306 + \frac{3}{2}(0.678091\eta + 0.26819\eta^2 + 0.0556078333\eta^3 \\ + 0.005151675\eta^4 - 0.001202982\eta^5 + 0.00008173300002\eta^6), \\ (0 < \eta \leq 3),$$

$$F_{3/2}(\eta) = 10.35369867 + \frac{3}{2}(-6.168596892 + 0.757064709\eta + 0.1961444\eta^2 \\ + 0.0901841666\eta^3 - 0.0042073325\eta^4 + 0.00016516628\eta^5 \\ - 0.000003032951667\eta^6), \quad (3 < \eta \leq 10),$$

$$F_{3/2}(\eta) = 134.270211 - (134.2701846 + 0.4\eta^{5/2} + 2.467401\eta^{1/2} \\ - 0.7102746\eta^{-3/2} - 2.771862428\eta^{-7/2} - 44.13000364\eta^{-11/2} \\ - 1,641.825466\eta^{-15/2}), \quad (10 < \eta \leq 10^5),$$

$$F_{3/2}(\eta) = \frac{2}{5}\eta^{5/2}, \quad (10^5 < \eta).$$

For additional information and other approximations see [McDougall and Stoner \(1938\)](#).

C.3.2 Relativistic Fermi Integrals

$$F_k^{(R)}(\eta, T') = \frac{1}{\Gamma(k+1)} \int_0^\infty \frac{x^k (1 + T'x)(1 + \frac{1}{2}T'x)^{1/2}}{e^{x-\eta} + 1} dx,$$

where $T' \equiv kT/(mc^2)$.

$$\lim_{T' \rightarrow 0} F_k^{(R)}(\eta, T') = F_k(\eta),$$

C.4 Hermite Polynomials

The Hermite differential equation is

$$\frac{d^2 w}{dx^2} - 2x \frac{dw}{dx} + 2nw = 0.$$

The Hermite polynomials are given by Rodrigues' formula

$$H_n(x) = (-1)^n e^{x^2} \frac{d^n}{dx^n} e^{-x^2},$$

$$H_n(x) = n! \sum_{k=0}^{n/2} \frac{(-1)^k}{k!(n-2k)!} (2x)^{n-2k}.$$

In particular:

$$\begin{aligned} H_0(x) &= 1, \\ H_1(x) &= 2x, \\ H_2(x) &= 4x^2 - 2, \\ H_3(x) &= 8x^3 - 12x, \\ H_4(x) &= 16x^4 - 48x^2 + 12, \\ H_5(x) &= 32x^5 - 160x^3 + 120x. \end{aligned}$$

For additional information see [Morse and Feshbach \(1953\)](#).

C.5 Generalized Laguerre Polynomials

$$L_n^\lambda(z) = \sum_{k=0}^n \frac{(-1)^k}{k!} \binom{n+\lambda}{n-k} z^k.$$

In particular:

$$L_0^\lambda(z) = 1,$$

$$L_1^\lambda(z) = -z + \lambda + 1,$$

$$L_2^\lambda(z) = \frac{z^2}{2} - (\lambda + 2)z + \frac{(\lambda + 2)(\lambda + 1)}{2},$$

$$L_3^\lambda(z) = -\frac{z^3}{6} + \frac{(\lambda + 3)z^2}{2} - \frac{(\lambda + 2)(\lambda + 3)z}{2} + \frac{(\lambda + 1)(\lambda + 2)(\lambda + 3)}{6}.$$

C.6 Legendre Polynomials

C.6.1 Unassociated Legendre Polynomials

The Legendre differential equation

$$(1 - x^2) \frac{d^2 w}{dx^2} - 2x \frac{dw}{dx} + l(l + 1)w = 0,$$

or

$$\frac{d}{dx} \left[(1 - x^2) \frac{dw}{dx} \right] + l(l + 1)w = 0,$$

where l is real, is a special case of the Associated Legendre equation (see Sect. C.6.2) with $m = 0$. Legendre polynomials are given by Rodrigues' formula

$$P_n(x) = \frac{1}{2^n n!} \frac{d^n}{dx^n} (x^2 - 1)^n.$$

$$\begin{aligned} P_n(z) &= \frac{1}{2^n} \sum_{k=0}^{n/2} (-1)^k \binom{n}{k} \binom{2n-2k}{n} z^{n-2k} \\ &= \frac{1}{2^{2n}} \sum_{k=0}^n \binom{2k}{n} \binom{2n-2k}{n-k} \cos[(n-2k) \arccos z], \quad (l \text{ is integer}). \end{aligned}$$

In particular:

$$P_0(x) = 1,$$

$$P_1(x) = x,$$

$$P_2(x) = \frac{1}{2} \cdot (3x^2 - 1),$$

$$P_3(x) = \frac{1}{2}(5x^3 - 3x),$$

$$P_4(x) = \frac{1}{8}(35x^4 - 30x^2 + 3),$$

$$P_5(x) = \frac{1}{8}(63x^5 - 70x^3 + 15x),$$

$$P_6(x) = \frac{1}{16}(231x^6 - 315x^4 + 105x^2 - 5).$$

For $x = \cos \theta$:

$$P_0(\cos \theta) = 1,$$

$$P_1(\cos \theta) = \cos \theta,$$

$$P_2(\cos \theta) = \frac{1}{4}[3 \cos(2\theta) + 1],$$

$$P_3(\cos \theta) = \frac{1}{8}[5 \cos(3\theta) + 3 \cos \theta],$$

$$P_4(\cos \theta) = \frac{1}{64}[35 \cos(4\theta) + 20 \cos(2\theta) + 9],$$

$$P_5(\cos \theta) = \frac{1}{128}[63 \cos(5\theta) + 35 \cos(3\theta) + 30 \cos \theta],$$

$$P_6(\cos \theta) = \frac{1}{512}[231 \cos(6\theta) + 126 \cos(4\theta) + 105 \cos(2\theta) + 50].$$

C.6.2 Associated Legendre Polynomials

The Legendre associated differential equation is

$$(1 - x^2) \frac{d^2 w}{dx^2} - 2x \frac{dw}{dx} + \left[l(l - 1) + \frac{m^2}{1 - x^2} \right] w = 0.$$

Associated Legendre functions of the first kind are given by

$$\begin{aligned}
 P_l^m(x) &= (1-x^2)^{m/2} \frac{d^m}{dx^m} P_l(x) \\
 &= \frac{(1-x^2)^{m/2}}{2^l l!} \frac{d^{m+l}}{dx^{m+l}} (x^2-1)^l. \\
 P_l^{-m}(x) &= (-1)^m \frac{(l-m)!}{(l+m)!} P_l^m(x).
 \end{aligned}$$

In particular:

$$\begin{aligned}
 P_0^0(x) &= 1, \\
 P_1^0(x) &= x, \\
 P_1^1(x) &= -(1-x^2)^{1/2}, \\
 P_2^0(x) &= \frac{1}{2}(3x^2-1), \\
 P_2^1(x) &= -3x(1-x^2)^{1/2}, \\
 P_2^2(x) &= 3(1-x^2), \\
 P_3^0(x) &= \frac{1}{2}x(5x^2-3), \\
 P_3^1(x) &= -\frac{3}{2}(1-5x^2)(1-x^2)^{1/2}, \\
 P_3^2(x) &= 15x(1-x^2), \\
 P_3^3(x) &= -15(1-x^2)^{3/2},
 \end{aligned}$$

For $x = \cos \theta$:

$$\begin{aligned}
 P_0^0(\cos \theta) &= 1, \\
 P_1^0(\cos \theta) &= \cos \theta, \\
 P_1^1(\cos \theta) &= -\sin \theta, \\
 P_2^0(\cos \theta) &= \frac{1}{2}(3 \cos^2 \theta - 1), \\
 P_2^1(\cos \theta) &= -3 \cos \theta \sin \theta, \\
 P_2^2(\cos \theta) &= 3 \sin^2 \theta, \\
 P_3^0(\cos \theta) &= \frac{1}{2}(5 \cos^3 \theta - 3 \cos \theta),
 \end{aligned}$$

$$P_3^1(\cos \theta) = -\frac{3}{2}(5 \cos^2 - 1) \sin \theta,$$

$$P_3^2(\cos \theta) = 15 \cos \theta \sin^2 \theta,$$

$$P_3^3(\cos \theta) = -15 \sin^3 \theta,$$

C.7 Coulomb Corrections to Pressure and Entropy of an Ideal Gas of Ions

Semi-empirical fits to data by [Brush et al. \(1966\)](#) for the Coulomb correction (see [van Horn 1971](#)).

$$\frac{(P - P_0)V}{NkT} = -0.113\Gamma^{3/2} \left[\frac{1}{(1 + 0.142\Gamma)^{1/2}} + \frac{1.54}{(1 + 0.575\Gamma)^{3/2}} \right],$$

$$\frac{(S - S_0)}{Nk} = -\ln \left[1 + \frac{\Gamma^{3/2}}{2\sqrt{3}} \left(0.015 + \frac{0.585}{1 + \Gamma^{1/2}} + \frac{0.400}{1 + 1.308\Gamma^{3/2}} \right) \right],$$

P_0 and S_0 are the pressure and entropy of a perfect gas of N ions in a box of volume V at temperature T . $\Gamma = (Ze)^2/(4\pi\epsilon_0 r_0 kT)$ is the ratio of the Coulomb interaction energy relative to the thermal energy kT , and r_0 is the ion sphere radius.

C.8 Spherical Harmonics

C.8.1 Scalar Spherical Harmonics

$$Y_l^m(\theta, \phi) = (-1)^m \left[\frac{(2l+1)(l-m)!}{4\pi(l+m)!} \right]^{1/2} P_l^m(\cos \theta) e^{im\phi}, \quad (-l < m < l).$$

$$Y_{l'}^{m'*}(\theta, \phi) = (-1)^m Y_l^{-m}(\theta, \phi),$$

where superscript * indicates the complex conjugate.

$$\int_{\phi=0}^{2\pi} \int_{\theta=0}^{\pi} Y_{l'}^{m'*}(\theta, \phi) Y_l^m(\theta, \phi) \sin \theta d\theta d\phi = \delta_{ll'} \delta_{mm'}.$$

In particular:

$$\begin{aligned}
 Y_0^0(\theta, \phi) &= \sqrt{\frac{1}{4\pi}}, \\
 Y_1^0(\theta, \phi) &= \sqrt{\frac{3}{4\pi}} \cos \theta, \\
 Y_1^{\pm 1}(\theta, \phi) &= \mp \sqrt{\frac{3}{8\pi}} \sin \theta e^{\pm i\phi}, \\
 Y_2^0(\theta, \phi) &= \frac{1}{4} \sqrt{\frac{5}{\pi}} (3 \cos^2 \theta - 1), \\
 Y_2^{\pm 1}(\theta, \phi) &= \mp \sqrt{\frac{15}{8\pi}} \sin \theta \cos \theta e^{\pm i\phi}, \\
 Y_2^{\pm 2}(\theta, \phi) &= \frac{1}{4} \sqrt{\frac{15}{2\pi}} \sin^2 \theta e^{\pm 2i\phi}.
 \end{aligned}$$

C.8.2 Vector Spherical Harmonics

$$\mathbf{M}_{\text{even}, mn}^{(z)} = \pm \frac{m \cos(m\phi)}{\sin \theta \sin(m\phi)} P_n^m(\cos \theta) z_n(\rho) \hat{\mathbf{e}}_\theta - \frac{\sin(m\phi)}{\cos(m\phi)} \frac{dP_n^m(\cos \theta)}{d\theta} z_n(\rho) \hat{\mathbf{e}}_\phi,$$

$$\begin{aligned}
 \mathbf{N}_{\text{odd}, mn}^{(z)} &= \frac{z_n(\rho) \cos(m\phi)}{\rho \sin(m\phi)} n(n+1) P_n^m(\cos \theta) \hat{\mathbf{e}}_r \\
 &+ \frac{\cos(m\phi)}{\sin(m\phi)} \frac{dP_n^m(\cos \theta)}{d\theta} \frac{1}{\rho} \frac{d[\rho z_n(\rho)]}{d\rho} \hat{\mathbf{e}}_\theta \\
 &\mp \frac{m \sin(m\phi)}{m \cos(m\phi)} \frac{P_n^m(\cos \theta)}{\sin \theta} \frac{1}{\rho} \frac{d[\rho z_n(\rho)]}{d\rho} \hat{\mathbf{e}}_\phi,
 \end{aligned}$$

here z_n stands for any one of four spherical Bessel functions of the first kind j_n , the second kind y_n , or the third kind (Hankel functions) $h_n^{(1)}$, or $h_n^{(2)}$. $\hat{\mathbf{e}}_\theta$ and $\hat{\mathbf{e}}_\phi$ are unit vectors.

Appendix D

Units, Conversion Factors, and Fundamental Physical Constants for Opacities

We present units, conversion factors, and some fundamental physical constants that are useful for opacity calculations. Many of these and others can be found on the Web at <http://physics.nist.gov/cgi-bin/cuu>. Updates on physical constants can be found at <http://physics.nist.gov/cuu/Constants/Index.html>.

Atomic mass constant:

$$10^{-3} \times \frac{M(^{12}\text{C})}{12} \frac{1}{N_o} = 1.660\,538\,86 \times 10^{-27} \text{ kg}$$

$$10^{-3} \times \frac{M(^{12}\text{C})}{12} \frac{1}{N_o} c^2 = 1.492\,417\,90 \times 10^{-10} \text{ J} = 931.494\,043 \text{ MeV}$$

Atomic unit of length (first Bohr radius of hydrogen), area, and volume:

$$a_o = \frac{\epsilon_o h^2}{\pi m e^2} = \frac{\alpha_o}{4\pi R_\infty} = 5.291\,772\,09 \times 10^{-11} \text{ m}$$

$$\pi a_o^2 = 8.797\,355\,37 \times 10^{-21} \text{ m}^2$$

$$\frac{4\pi}{3} a_o^3 = 6.207\,146\,61 \times 10^{-31} \text{ m}^3$$

Atomic unit of time:

$$\tau_o = (4\pi R_\infty c)^{-1} = 2.418\,884\,326 \times 10^{-17} \text{ s}$$

Avogadro constant:

$$N_o = 6.022\,141\,5 \times 10^{23} \text{ atoms (g-mol)}^{-1}$$

Bohr radius: see atomic unit of length

Boltzmann constant:

$$k = 1.380\,650\,5 \times 10^{-23} \text{ J K}^{-1} = 8.617\,343 \times 10^{-5} \text{ eV K}^{-1}$$

$$\frac{k}{h} = 2.083\,664\,4 \times 10^{10} \text{ Hz K}^{-1}$$

$$\frac{k}{hc} = 69.503\,564 \text{ m}^{-1} \text{ K}^{-1}$$

Classical electron radius:

$$r_{ce} = \frac{e^2}{4\pi\epsilon_0 mc^2} = \alpha_0^2 a_0 = 2.817\,940\,29 \times 10^{-15} \text{ m}$$

Compton wavelength:

$$\lambda_C = \frac{h}{mc} = 2\pi\alpha_0 a_0 = \frac{\alpha^2}{2R_\infty} = 2.426\,310\,22 \times 10^{-12} \text{ m}$$

Einstein B_{lu} coefficient for radiation density per unit frequency interval at the frequency of the absorption line in terms of the oscillator strength:

$$B_{lu} = \frac{e^2}{4h\nu m\epsilon_0} f_{lu}$$

Conversion factors to Einstein B_{lu} coefficient for radiation density per unit frequency interval at the frequency of the absorption line [$\text{s}^{-1}(\text{Jm}^{-3}\text{Hz}^{-1})^{-1}$] from:

B_{lu} coefficient for radiation density per unit wavelength interval at the wavelength of the absorption line [$\text{s}^{-1}(\text{Jm}^{-3}\text{m}^{-1})^{-1}$]: ν^2/c ,

B_{lu} coefficient for radiance per unit wavelength interval at the wavelength of the absorption line [$\text{s}^{-1}(\text{Wm}^{-2}\text{sr}^{-1}\text{m}^{-1})^{-1}$]: $\nu^2/(4\pi)$,

B_{lu} coefficient for radiance per unit frequency interval at the frequency of the absorption line [$\text{s}^{-1}(\text{Wm}^{-2}\text{sr}^{-1}\text{Hz}^{-1})^{-1}$]: $c/(4\pi)$.

Electric charge:

$$1 \text{ C} = 1 \text{ F} \times 1 \text{ V}$$

Electric dipole moment:

$$1 \text{ debye} \equiv 10^{-18} \text{ esu cm} = 3.335\,64 \times 10^{-30} \text{ C m}$$

$$ea_0 = 8.478\,353\,07 \times 10^{-30} \text{ C m}$$

$$e \times 1 \text{ \AA} = 1.602\,176\,53 \times 10^{-29} \text{ C m}$$

$$e \times 1 \text{ nm} = 1.602\,176\,53 \times 10^{-28} \text{ C m}$$

Electric quadrupole moment:

$$ea_0^2 = 4.486\,551\,23 \times 10^{-40} \text{ C m}^2$$

Electric potential:

$$1 \text{ V} = 1 \text{ J/C}$$

Electron mass:

$$m = 9.109\,382 \times 10^{-31} \text{ kg} = 5.485\,799 \times 10^{-4} \text{ amu}$$

$$(N_0 10^3 m)^{-1} = 1,822.89 \text{ amu/electron}$$

Electron mass energy equivalent:

$$mc^2 = 8.187\,104 \times 10^{-14} \text{ J} = 0.510\,998\,9 \text{ MeV} = 2/\alpha_0^2 \text{ ryd}$$

Electron plasma frequency:

$$\nu_p = \left(\frac{N_e e^2}{4\pi^2 m \epsilon_0} \right)^{1/2} = 8.978\,662\,8 N_e^{1/2} \text{ Hz}$$

Elementary charge:

$$e = 1.602\,176\,5 \times 10^{-19} \text{ C}$$

$$\frac{e^2}{4\pi \epsilon_0} = 2.307\,077\,2 \times 10^{-28} \text{ J m}$$

$$\frac{e^2}{4\pi \epsilon_0} \cdot \frac{\pi}{mc} = 2.654\,009 \times 10^{-6} \text{ m}^2 \text{ s}$$

Energy (see also Hartree and Rydberg units):

$$1 \text{ J} = 6.241\,509\,48 \times 10^{18} \text{ eV} = 4.587\,425\,13 \times 10^{17} \text{ ryd}$$

$$1 \text{ F} = 1 \text{ C/V}$$

$$1 \text{ C V} = 1 \text{ J}$$

$$1 \text{ eV} = 1.602\,176\,53 \times 10^{-19} \text{ J}$$

$$1 \text{ eV}/ch = 8.065\,544\,45 \times 10^5 \text{ m}^{-1}$$

$$1 \text{ ryd} = 2.179\,872\,1 \times 10^{-18} \text{ J}$$

$$1 \text{ ryd}/ch = 1.097\,367\,7 \times 10^7 \text{ m}^{-1}$$

$$1 \text{ eV}/k = 1.160\,450\,5 \times 10^4 \text{ K}$$

Fine structure constant:

$$\alpha_0 = \frac{e^2}{2\varepsilon_0 hc} = 7.297\,352\,54 \times 10^{-3}$$

$$\alpha_0^{-1} = 137.035\,999\,7$$

Hartree unit of energy (= 2 Rydberg units of energy):

$$2R_\infty hc = \frac{e^2}{4\pi\varepsilon_0 a_0} = 4.359\,744\,17 \times 10^{-18} \text{ J} = 27.211\,384\,5 \text{ eV}$$

Ion plasma frequency

$$\nu_{\text{pi}} = \left(\frac{N_i (Ze)^2}{4\pi^2 \varepsilon_0 m_i} \right)^{1/2} = 5.160\,682 \times 10^{12} \frac{Z}{A} \rho^{1/2} \text{ Hz}$$

Loschmidt constant:

$$N_L = \frac{N_0}{V_M} = 2.686\,777 \times 10^{25} \text{ m}^{-3} \quad (\text{at } T = 273.15 \text{ K, } P = 101.325 \text{ kPa}).$$

The Loschmidt constant must be used with care, because sometimes (e.g., see CODATA definitions) it is defined at $P = 100 \text{ kPa}$. In that case $N_L = 2.651\,646 \times 10^{25} \text{ m}^{-3}$. In this book it is always defined for $P = 1 \text{ atm}$.

Molar gas constant:

$$R_0 = kN_0 = 8.314\,472 \text{ J (g-mol)}^{-1} \text{ K}^{-1}$$

Molar volume of ideal gas:

$$\frac{R_0 T}{P} = V_M = 22.413\,996 \times 10^{-3} \text{ m}^3 (\text{g-mol})^{-1} \quad (\text{at } T = 273.15 \text{ K,}$$

$$P = 101.325 \text{ kPa})$$

Permeability of vacuum:

$$\mu_0 = 4\pi \times 10^{-7} \text{ N A}^{-2} = 1.256\,637\,061\,4 \dots \times 10^{-6} \text{ N A}^{-2} \quad (\text{exact})$$

Permittivity of vacuum:

$$\varepsilon_0 = 1/(4\pi \times 10^{-7} c^2) = 8.854\,187\,817 \dots \times 10^{-12} \text{ F m}^{-1} \quad (\text{exact})$$

$$\frac{1}{4\pi\varepsilon_0} = 8.987\,551\,178\,8 \times 10^9 \text{ V m C}^{-1}$$

Planck constant:

$$h = 6.626\,068\,96 \times 10^{-34} \text{ J s} = 4.135\,667\,33 \times 10^{-15} \text{ eV s}$$

$$hc = 1.986\,455\,44 \times 10^{-25} \text{ J m}$$

Plasma frequency see: electron or ion plasma frequency

Proton mass:

$$1.672\,621\,71 \times 10^{-27} \text{ kg}$$

Radiation constant (first):

$$c_1 = 2\pi hc^2 = 3.741\,771\,38 \times 10^{-16} \text{ W m}^2$$

Radiation constant (second):

$$c_2 = \frac{hc}{k} = 1.438\,775\,2 \times 10^{-2} \text{ m K}$$

Rydberg constant for infinite mass:

$$R_\infty = \left(\frac{e^2}{4\pi\epsilon_0} \right)^2 \frac{2\pi^2 m}{h^2 c} = 1.097\,373\,156\,852\,5 \times 10^7 \text{ m}^{-1}$$

Rydberg frequency:

$$R_\infty c = 3.289\,841\,960\,360 \times 10^{15} \text{ Hz}$$

Rydberg unit of energy:

$$R_\infty hc = \frac{e^2}{8\pi\epsilon_0 a_0} = 2.179\,872\,09 \times 10^{-18} \text{ J} = 13.605\,692\,3 \text{ eV} \equiv 1 \text{ ryd}$$

Speed of light in vacuum:

$$c = 2.997\,924\,58 \times 10^8 \text{ m s}^{-1} \quad (\text{exact})$$

Stefan constant:

$$a_S = 4 \frac{\sigma}{c} = 7.565\,767\,4 \times 10^{-16} \text{ J m}^{-3} \text{ K}^{-4} = 4.722\,180\,9 \times 10^3 \text{ eV m}^{-3} \text{ K}^{-4}$$

Stefan-Boltzmann constant:

$$\sigma = \frac{2\pi^5 k^4}{15h^3 c^2} = 5.670\,400 \times 10^{-8} \text{ W m}^{-2} \text{ K}^{-4} = 3.539\,186 \times 10^{11} \text{ eV m}^{-2} \text{ K}^{-4} \text{ s}^{-1}$$

Thomson cross section:

$$\sigma_{\text{T}} = \frac{8\pi}{3} r_{\text{ce}}^2 = 6.652\,458\,73 \times 10^{-29} \text{ m}^2$$

Conversion of dimensionless polarizability ($\alpha_{\text{p}}/4\pi\epsilon_0 a_0^3$) to polarizability in F m^2 (or $\text{C m}^2 \text{ V}^{-1}$):

$$4\pi\epsilon_0 a_0^3 = 1.648\,777\,274 \times 10^{-41} \text{ F m}^2$$

Conversion of polarizability in F m^2 to molar polarizability in $\text{m}^3/(\text{g-mol})$:

$$\frac{N_0}{3\epsilon_0} = 2.267\,153\,7 \times 10^{34} \text{ m}^3/\text{F m}^2(\text{g-mol})$$

Conversion of polarizability in SI units to cgs units in cm^3 :

$$\alpha_{\text{p}}[\text{cm}^3] = \frac{100h}{4\pi\epsilon_0} \frac{\alpha}{h} [\text{Hz}/(\text{V}/\text{cm})^2] = 5.955\,213\,79 \times 10^{-22} \frac{\alpha}{h} [\text{Hz}/(\text{V}/\text{cm})^2]$$

Conversion from wavelength in vacuum to energy:

$$E = \frac{hc}{\lambda_{\text{vac}}} = 1.986\,455\,44 \times 10^{-16} \text{ J } (\lambda_{\text{vac}} \text{ in nm})$$

$$E = \frac{hc}{\lambda_{\text{vac}}} = \frac{1,239.8480}{\lambda_{\text{vac}}} \text{ eV } (\lambda_{\text{vac}} \text{ in nm})$$

$$E = \frac{hc}{\lambda_{\text{vac}}} = \frac{91.127\,164}{\lambda_{\text{vac}}} \text{ ryd } (\lambda_{\text{vac}} \text{ in nm})$$

Appendix E

Some Relevant Websites

The intent here is not to give an exhaustive list of websites, but to give a short list of some websites relevant for opacities.

Atomic Physics Codes:

Codes based on R. D. Cowan's *The Theory of Atomic Structure and Spectra*:
<http://www.tcd.ie/Physics/people/Cormac.McGuinness/Cowan/>

Atomic, Molecular, and Electron Collisions:

<http://aphysics2.lanl.gov/tempweb/lanl/>
<http://www.ruf.rice.edu/~atmol/index.html>

Databases:

<http://www.nist.gov/srd/>
High resolution TRANsmission (HITRAN) database:
<http://www.cfa.harvard.edu/HITRAN>

Electron Conductivity for Stellar Plasmas:

<http://www.ioffe.rssi.ru/astro/conduct/conduct.html>

Electron Impact Cross Sections for Ionization and Excitation:

<http://physics.nist.gov/PhysRefData/Ionization/Xsection.html>
International Atomic Energy Agency:
<http://www-amdis.iaea.org/databases.html>
National Institute for Fusion Science, Data and Planning Center (Nagoya, Japan):
<http://dbshino.nifs.ac.jp>

EOS and Associated Data for Materials in the Sesame Library:

<http://t1web.lanl.gov>
<http://www.lanl.gov/orgs/t/t1.shtml>
First-time users click on "Sesame registration."

Gaussian Basis Sets are Described in GAUSSIAN-98 User's Guide:

<http://www.gaussian.com/>

Kinetics Databases:

<http://kinetics.nist.gov/kinetics/>

<http://kinetics.nist.gov/solution/>

<http://www.udfa.net>

Rate Coefficients for Radiative Processes:

Solar and blackbody radiation:

<http://phidrates.space.swri.edu>

Photodissociation and photoionization of astrophysically relevant molecules:

<http://home.strw.leidenuniv.nl/~ewine/photo/>

Materials Libraries:

<http://www.codata.org/>

Click on “Databases”, then click on “International Register of Materials Database Managers.”

Opacities:

Opacity (astrophysical and general): <http://aphysics2.lanl.gov/opacity/lanl/>

Opacity Project (TOPbase):

<http://cdsweb.u-strasbg.fr/topbase.html>

Livermore (OPLIB):

<http://www-phys.llnl.gov/Research/OPAL/opal.html>

<http://www.cita.utoronto.ca/~boothroy/kappa.html>

<http://www.physast.uga.edu/ugamop>

Low-temperature opacities:

<http://webs.wichita.edu/physics/opacity/>

Programs and data for CIA opacities:

<http://www.astro.ku.dk/~aborysow/programs/>

Physics Data Libraries:

<http://www.fas.harvard.edu/~planets/planetarylinks.html>

Thermodynamic Data:

<http://kinetics.nist.gov/janaf>

Chemical Equilibrium:

<http://www.grc.nasa.gov/WWW/CEAWeb/ceaRequestForm.htm>

Units, Conversion Factors, and Fundamental Physical Constants:

<http://physics.nist.gov/cuu/Constants/index.html>

<http://spintronics.inha.ac.kr/unit.pdf>

<http://newton.ex.ac.uk/research/qsystems/collabs/constants.html>

<http://qft.iqfr.csic.es/PhysRefData/codata86/codata86.html>

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