



## ONE-POT MANNICH BASE SYNTHESIS USING TASK SPECIFIC PROTIC IONIC LIQUIDS

(Sintesis Bes Mannich Tunggal Menggunakan Cecair Ionik Terkhusus)

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Received: 21 August 2016; Accepted: 27 July 2017

### Abstract

Three-component Mannich reaction of ketones, aromatic aldehydes and aromatic amines was catalyzed by four Bronsted acidic ionic liquids comprising of iodide and borate at room temperature. Ionic liquid have been used as catalyst and solvent to produce some Mannich bases in high yield (75%) and shorter reaction time (20 minutes). Work up has been facilitated by simple extraction with water to recover ionic liquid for recycling up to four times without any significant loss in activity.

**Keywords:** Mannich bases, Bronsted acidic ionic liquids, 1-methylimidazole, 1,3-propane sultone, 1,4-butane sultone

### Abstrak

Tindak balas tiga komponen Mannich iaitu keton, aldehid aromatik dan amina aromatik telah dimungkinkan oleh empat cecair ionik Bronsted berasid yang terdiri daripada iodida dan borat pada suhu bilik. Cecair ionik telah digunakan sebagai pemangkin dan pelarut untuk menghasilkan beberapa bes Mannich dengan hasil yang tinggi (75%) dan masa tindak balas yang lebih pendek (20 minit). Kerja pengekstrakan telah dipermudah dengan menggunakan air dan boleh dikitar semula sehingga empat kali tanpa sebarang kehilangan aktiviti yang penting.

**Kata kunci:** bes Mannich, asid Bronsted cecair ionik, 1-metilimidazol, 1,3-propana sulton, 1,4-butana sulton

### Introduction

Multicomponent reactions have been proven to be an economic and facile approach for building of complex molecules in one synthetic step. A remarkable diversity can be achieved by varying the starting materials. Such appealing reactions offer an elegant course for synthesis of carbon-carbon and carbon-hetero atom linkages with high atom economy, simple methodology and high selectivity in one-pot procedure [1]. Mannich reaction is a fundamental method for a formation of  $\beta$ -amino carbonyl compounds of synthetic intermediates for natural product synthesis (alkaloids and polyketides) and in pharmaceutical industry [2]. The conventional catalysts for the synthesis of classical Mannich bases are mainly organic and mineral acids such as acetic acid [3], proline [4], p-dodecylbenzenesulfonic acid [5] and lewis acids [6]. These reactions conditions often lead to problems like difficulty in product separation, harsh reaction conditions, long reaction durations, toxicity and pollution [7]. Further development lead to the usage of other catalysts such as pTSA, BDMS,  $\text{HClO}_4\text{-SiO}_2$ ,  $\text{TMSCl}$ , Amberlyst-15, carbon based solid acid,  $\text{AuCl}_3\text{-PPh}_3$ ,  $\text{InCl}_3$ , silica supported sulphuric acid,  $\text{NbCl}_5$ , bismuth (III) chloride, sulfamic acid, boric acid, sucrose charsulfonic acid,  $\text{Zn}(\text{OTf})_2$ , etc. [8]. The drawbacks include: large amount of

catalyst, expensive catalyst, high waste production, long reaction duration and ultimately low yield. [9]. Lanthanide triflate in acetonitrile or dichloromethane has also been reported for Mannich reaction catalysis [10].

In contribution to green chemistry, many studies reported on employing traditional volatile and flammable solvents to benign alternatives such as non-volatile solvent, supercritical system, aqueous media or solventless conditions. In this context, Mannich type reactions have also been reported in water [10]. In current years, ionic liquids (ILs) have been used as an environmentally benign solvents, efficient catalysts and a promising substitutes for volatile organic solvents and traditionally used acid/ base catalysts in construction of organic moieties [11, 12]. Use of organic solvents in reaction mixture usually leads to serious safety issues such as volatility, flammability and toxicity [12, 13]. The use of ionic liquids as reaction media offers a convenient solution to both solvent emission and catalytic recycling problems [10]. Beside high activity and selectivity, these versatile liquids are center of interest as they possess unique properties such as flexibility, high stability in air and water, easy separation and purification processes, reusability, non-volatility, non-flammability, noncorrosive nature, controlled immiscibility, low toxicity, low vapour pressure, low viscosities, high thermal and chemical stabilities [14-16]. Recently work on Mannich reaction is being carried out using ionic liquids [17, 18].

The present study emphasizes on employing ionic liquids as catalyst and solvents to overcome issues associated with traditional catalysts and organic solvents. In this approach, a green procedure has been proposed for synthesis of Mannich bases using acetophenone, aromatic aldehydes and anilines as starting materials. To the best of our knowledge, Mannich reaction catalyzed by ILs bearing imidazolium sulfone as cation; and iodide and borate as anions are unprecedented. The reaction proceeded at faster rate with higher yield of desired Mannich bases using the synthesized ionic liquids. The effect of cations and anions in addition to the influence of additional carbon atom on sulfone moiety was also been investigated.

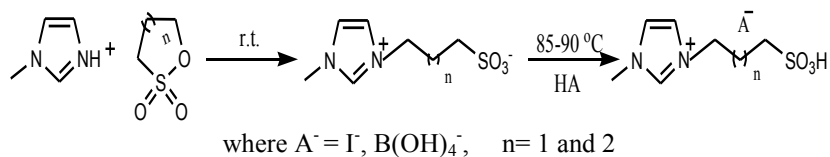
## Materials and Methods

### Chemicals

All the reacting molecules used such as 1,3-propane sultone, 1,4-butane sultone, *N*-methylimidazole, hydriodic acid, boric acid, acetonitrile, aromatic amines, aromatic aldehydes, acetophenone and cyclohexanone were purchased from Sigma Aldrich and used without any purification procedure. Melting points of the Mannich bases were determined in open capillaries and are uncorrected. NMR spectra were recorded on 400 MHz NMR Joel spectrometer using D<sub>2</sub>O and CDCl<sub>3</sub> as solvents. IR spectra were recorded in Nicolet FTIR spectrometer (400 – 4000 cm<sup>-1</sup>).

### Synthesis of ionic liquids

The task specific ionic liquids have been synthesized by using an analogous method as reported in literature [19-21] with slight modifications. 1,3-propane sultone and 1,4-butane sultone were reacted with *N*-methylimidazole in equimolar ratio affording the respective zwitter ions that were then converted to ILs by acidification with two protic acids (Scheme 1). The corresponding ILs were used further without any purification.

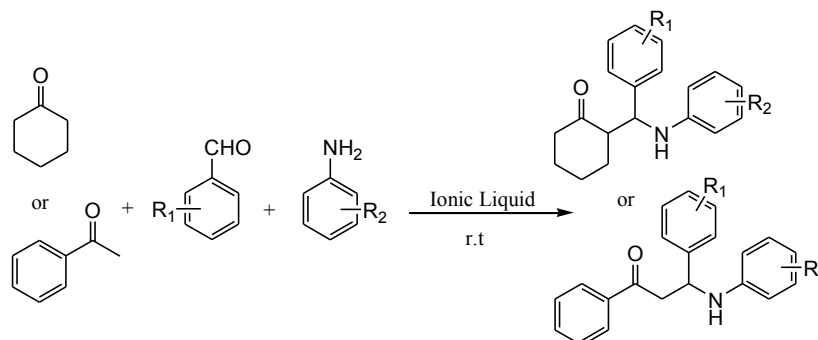


Scheme 1. Reaction scheme for IL synthesis

### Mannich reaction: A typical procedure

Cyclohexanone (10 mmol), benzaldehyde (10 mmol), aniline (10 mmol) and 5 mmol of catalyst were taken in round bottom flask at 25 °C (Scheme 2) [22]. The completion of reaction was determined using TLC. For work up, some amount of water was added and extraction with ethyl acetate was done. The excess solvent was evaporated via rotary evaporator to afford solid product which was recrystallized from acetonitrile and dried under vacuum until

constant weight was obtained. At the end, the aqueous layer containing IL was vacuum dried at 70 °C to reuse ionic liquid for future use. The synthesized compounds were then characterized using IR and NMR. The spectral analysis along with physical data (melting points) was consistent with those reported in literature.



Scheme 2. Mannich reaction catalyzed by imidazolium based ionic liquids (R<sub>1</sub> and R<sub>2</sub> are mentioned in Table 4)

### Results and Discussion

The ionic liquids were synthesized by the reaction of methyl imidazole with 1,3-propane sultone and 1,4-butane sultone for preparation of Mannich bases. Spectral analysis confirmed the structure of synthesized ILs. these synthesized ILs were used to catalyze Mannich reaction using acetophenone/ cyclohexanone (10 mmol), benzaldehyde (10 mmol) and aniline (10 mmol) as model substrates at room temperature (Scheme 2).

The reaction mixture gradually turned to highly viscous material within 30 minutes. The reaction work up was done by addition of a small amount of water followed by extraction with ethyl acetate and finally evaporation of organic solvent. For purification, the synthesized Mannich bases were recrystallized from acetonitrile and the aqueous layer was further dried for recycling ionic liquid. Among the four protic ILs 75% of Mannich base was obtained as highest yield by 1-methyl-3-(3-sulfopropyl)-imidazolium borate ([mimps]HBO) in 20 minutes (Table 1).

Table 1. Effect of ionic liquid on yield of Mannich base<sup>a</sup>

Entry	Catalyst	Solvent Used	Reaction Time (min)	Isolated Yield (%)
01	-	-, EtOH <sup>b, c, d</sup>	120	NR <sup>a, b, c, d</sup>
02	[mimps]I	-, EtOH <sup>b, c, d</sup>	120	NR <sup>a, b, c, d</sup>
03	[mimbs]I	-, EtOH <sup>b, c, d</sup>	120	NR <sup>a, b, c, d</sup>
04	[mimps]HBO	-	20	75
05	[mimbs]HBO	-	30	69

<sup>a</sup> Reaction conditions: cyclohexanone (10 mmol), benzaldehyde (10 mmol), aniline (10 mmol), catalyst (5 mmol) r.t.

<sup>b</sup> yield in presence of solvent

<sup>c</sup> yield obtained at 60 °C

<sup>d</sup> yield obtained at 80 °C

The reaction with I<sup>-</sup> as anion, was also carried out with ethanol using same IL as catalyst using different temperatures but the results obtained were insignificant (Table 1). It was observed that IL based on propane sultone gave better yield than butane sultone based IL (entries 4 and 5 in Table 1). Furthermore it was also found that the IL formed from weak acid (H<sub>3</sub>BO<sub>3</sub>) gave better yield as compared to the IL which was formed from strong acid (HI).

This observation might be rationalized on basis of pKa, which depicts that the reaction is catalyzed by conjugate bases ( $I^-$  and  $B(OH)_4^-$ ) of the protic acids ( $HI$  and  $H_3BO_3$ ).

For determining the amount of IL used, the reaction of cyclohexanone (10 mmol), benzaldehyde (10 mmol), aniline (10 mmol) was catalyzed under variable amounts of [mimps]HBO at room temperature. It was found that the yields of the Mannich base was increased with the increase in the amount of [mimps]HBO (Table 2), where 5 mmol was found as an optimal amount.

Table 2. Effect of amount of ionic liquid on Mannich reaction

Entry	Ionic Liquid (mmol)	Reaction Time (min)	Isolated Yield (%)
01	0	24 h	0
02	1	30	66
03	3	30	71
04	5	30	75
05	7	30	74
06	9	30	75

Mannich reaction was also conducted in polar and non-polar organic solvents to check the effect of solvent on the yield of product formed for estimation of the best reaction conditions. Herein the reaction of cyclohexanone, benzaldehyde and aniline was selected as standard reaction. It was found that the reaction in which no solvent was used gave the highest yield of 75% among the solvents selected showing that the ionic liquid may act as dual-reagent without need of any external solvent (Table 3).

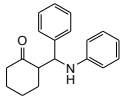
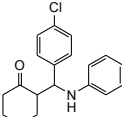
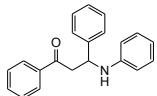
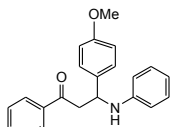
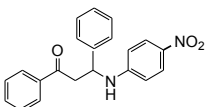
Table 3. Effect of ionic liquid/solvent on Mannich reaction<sup>a</sup>

Entry	Ionic Liquid/Solvent	Isolated Yield
01	[mimps]HBO	75
02	[mimps]HBO/H <sub>2</sub> O	70
03	[mimps]HBO/C <sub>2</sub> H <sub>5</sub> OH	69
04	[mimps]HBO/CH <sub>2</sub> Cl <sub>2</sub>	61
05	[mimps]HBO/C <sub>6</sub> H <sub>6</sub>	39

<sup>a</sup> Reaction conditions: cyclohexanone (10 mmol), benzaldehyde (10 mmol), aniline (10 mmol), [mimps]HBO (5 mmol) r.t.

The Mannich reaction was further investigated with different aromatic aldehydes and anilines using [mimps]HBO. It was observed that aromatic aldehyde with electron-donating group produced higher yield of Mannich base as compared to that of electron withdrawing group (Table 4). In case of aromatic anilines, both electron withdrawing and electron-donating groups have been reported to be equally valuable for Mannich reaction [7]. The synthesis of Mannich bases was confirmed by physical and spectroscopic data and was found in good agreement with the previous reports.

Table 4. Preparation of  $\alpha$ -amino ketones using [mimps]HBO as catalyst

Entry	R <sub>1</sub>	R <sub>2</sub>	H-Active Molecule	Product	Yield (%)	mp (°C) (lit. mp) [ref]
01	H	H	(CH <sub>2</sub> ) <sub>5</sub> CO		75	114 – 115 (115 – 116) [24]
02	4-chloro	H	(CH <sub>2</sub> ) <sub>5</sub> CO		73	133 – 134 (134 – 135) [25]
03	H	H	C <sub>6</sub> H <sub>5</sub> CO		75	168 – 169 (169 – 171) [26]
04	4-methoxy	H	C <sub>6</sub> H <sub>5</sub> CO		74	150 – 151 151 – 152 [27]
05	H	4-NO <sub>2</sub>	C <sub>6</sub> H <sub>5</sub> CO		72	184 – 186 183 – 185 [27]

Regarding the reusability of ionic liquid, it was found that the IL could be successfully reused up to four cycles without considerable loss of activity (Table 5).

Table 5. Reusability of [mimps]CH<sub>3</sub>COO for Mannich reaction

Entry	Cycle	% Yield
1	I	75
2	II	75
3	III	74
4	IV	74

#### Spectral analyses of Zwitter ions and ionic liquids: 1-methylimidazolium-3-(3-propylsulfonate)

<sup>1</sup>H NMR (400 MHz, D<sub>2</sub>O):  $\delta$  (ppm) 2.24 (2H, m), 2.73 (2H, t  $J=7.32$  Hz), 3.92 (3H, s), 4.24 (2H, t  $J=7.24$  Hz), 7.37 (1H, s), 7.45 (1H, s), 8.87 (1H, s); <sup>13</sup>C NMR (400 MHz, D<sub>2</sub>O):  $\delta$  (ppm) 25.5 (CH<sub>3</sub>), 38.7 (CH<sub>2</sub>), 48.2 (CH<sub>2</sub>), 49.5 (CH<sub>2</sub>), 124.1 (C=C), 126.3 (C=C), 139.1 (C=N); IR (cm<sup>-1</sup>): 3093 C-H stretching (alkene), 1654 C=N stretching (imine), 1456 CH<sub>2</sub> bending, 1178 and 1158 S=O stretching (sulfonate), 1034 S=O stretching (sulfoxide).

#### 1-methyl-3-(3-sulfopropyl)-imidazolium iodide

([mimps]I, Table 2, entry 2): <sup>1</sup>H NMR (400 MHz, D<sub>2</sub>O):  $\delta$  (ppm) 2.18 (2H, m), 2.78 (2H, t  $J=7.23$  Hz), 3.75 (3H, s), 4.22 (2H, t), 6.37 (1H, s), 6.45 (1H, s), 7.67 (1H, s); <sup>13</sup>C NMR (400 MHz, D<sub>2</sub>O):  $\delta$  (ppm) 22.9 (CH<sub>3</sub>), 37.9 (CH<sub>2</sub>), 50.4 (CH<sub>2</sub>), 56.8 (CH<sub>2</sub>), 122.9 (C=C), 123.1 (C=C), 137.2 (C=N); IR (cm<sup>-1</sup>): 3090 C-H stretching (alkene), 1685 C=N stretching (imine), 1560 C=C stretching (cyclic alkene), 1453 CH<sub>2</sub> bending, 1160 S=O stretching (sulfonate), 1020 S=O stretching (sulfoxide).

**1-methyl-3-(4-sulfobutyl)-imidazolium iodide**

([mimbs]I, entry 3):  $^1\text{H}$  NMR (400 MHz,  $\text{D}_2\text{O}$ ):  $\delta$  (ppm) 1.61 (2H, m), 1.88 (2H, m), 2.81 (2H, t  $J=6.98$  Hz), 3.73 (3H, s), 4.12 (2H, t  $J=6.86$  Hz), 7.32 (1H, s), 7.38 (1H, s), 8.61 (1H, s);  $^{13}\text{C}$  NMR (400 MHz,  $\text{D}_2\text{O}$ ):  $\delta$  (ppm) 21.3 ( $\text{CH}_3$ ), 28.2 ( $\text{CH}_2$ ), 37.7 ( $\text{CH}_2$ ), 52.1 ( $\text{CH}_2$ ), 57.6 ( $\text{CH}_2$ ), 122.6 (C=C), 122.8 (C=C), 137.4 (C=N); IR ( $\text{cm}^{-1}$ ): 3090 C-H stretching (alkene), 1685 C=N stretching (imine), 1560 C=C stretching (cyclic alkene), 1453  $\text{CH}_2$  bending, 1161 S=O stretching (sulfonate), 1020 S=O stretching (sulfoxide).

**1-methyl-3-(3-sulfopropyl)-imidazolium borate**

([mimps]HBO, entry 4):  $^1\text{H}$  NMR (400 MHz,  $\text{D}_2\text{O}$ ):  $\delta$  (ppm) 2.19 (2H, m), 2.78 (3H, s), 3.75 (2H, t  $J=7.24$  Hz), 4.23 (2H, t  $J=7.16$  Hz), 7.32 (1H, s), 7.38 (1H, s), 8.61 (1H, s);  $^{13}\text{C}$  NMR (400 MHz,  $\text{D}_2\text{O}$ ):  $\delta$  (ppm) 23.6 ( $\text{CH}_3$ ), 36.3 ( $\text{CH}_2$ ), 50.3 ( $\text{CH}_2$ ), 56.5 ( $\text{CH}_2$ ), 121.2 (C=C), 121.5 (C=C), 137.5 (C=N); IR ( $\text{cm}^{-1}$ ): 3159 C-H stretching (alkene), 1640 C=N stretching (imine), 1574 C=C stretching (cyclic alkene), 1392, 1163 S=O stretching (sulfonate), 1035 S=O stretching (sulfoxide).

**1-methyl-3-(4-sulfobutyl)-imidazolium borate**

([mimbs]HBO, entry 5):  $^1\text{H}$  NMR (400 MHz,  $\text{D}_2\text{O}$ ):  $\delta$  (ppm) 1.61 (2H, m), 1.91 (2H, m), 2.82 (2H, t  $J=7.24$  Hz), 3.78 (3H, s), 4.14 (2H, t  $J=7.31$  Hz), 7.35 (1H, s), 7.41 (1H, s), 8.61 (1H, s);  $^{13}\text{C}$  NMR (400 MHz,  $\text{D}_2\text{O}$ ):  $\delta$  (ppm) 20.03 ( $\text{CH}_3$ ), 27.2 ( $\text{CH}_2$ ), 36.9 ( $\text{CH}_2$ ), 47.9 ( $\text{CH}_2$ ), 49.2 ( $\text{CH}_2$ ), 121.8 (C=C), 122.6 (C=C), 136.3 (C=N); IR ( $\text{cm}^{-1}$ ): 3157 C-H stretching (alkene), 1641 C=N stretching (imine), 1575 C=C stretching (cyclic alkene), 1165 S=O stretching (sulfonate), 1038 S=O stretching (sulfoxide).

**Conclusion**

Present methodology deals with a useful and effective way for synthesis of Mannich bases at room temperature in short reaction duration (20 minutes) and good yield (75%) using ILs as dual reagents. The reported synthetic method can be expected as an environmentally green procedure that might give valuable organic transformations by substituting conventional volatile, toxic and flammable organic solvents with ILs. Simple experimental and work up conditions along with reusability of IL can contribute a new method of green chemistry concept.

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